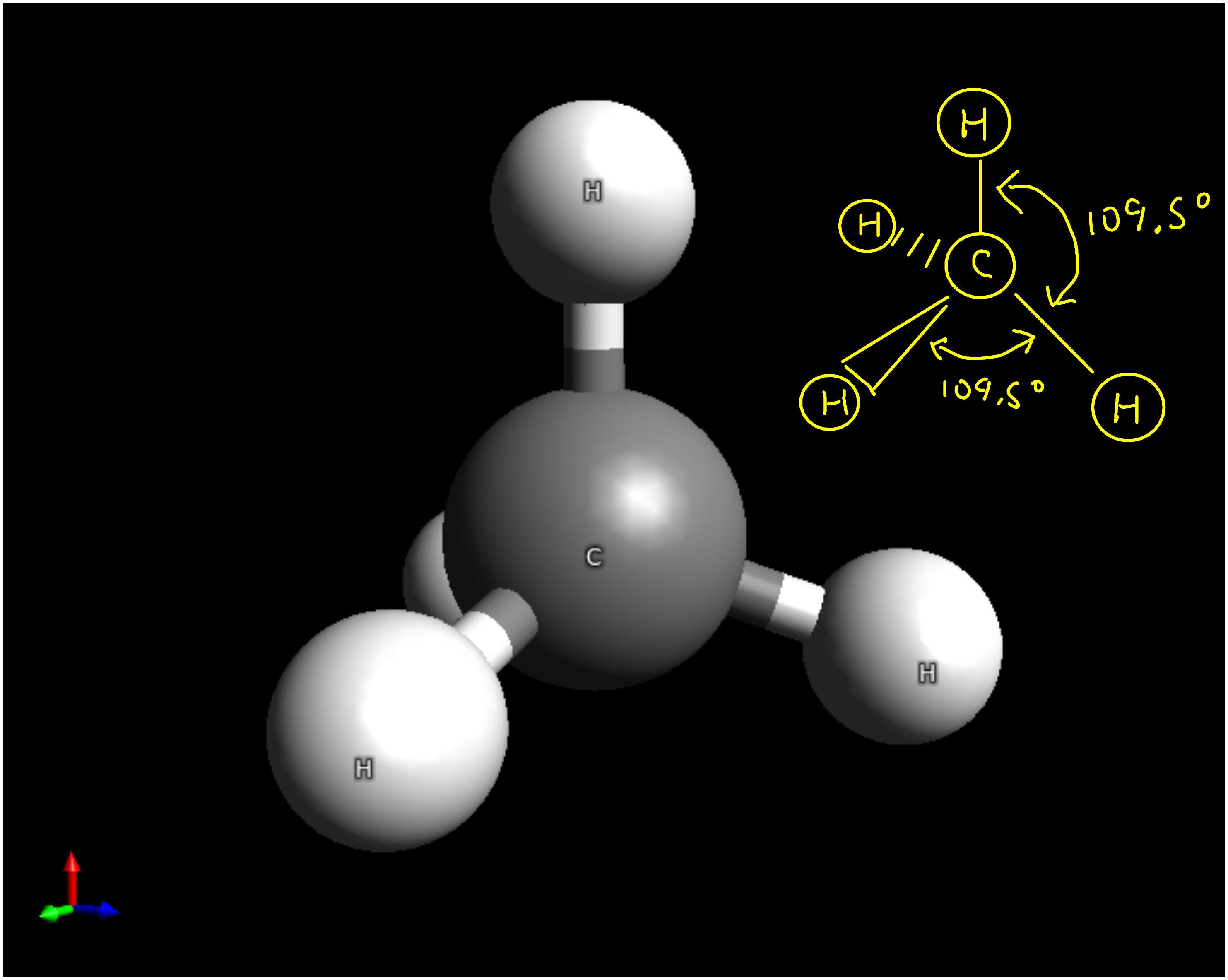
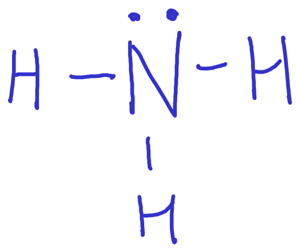
Here's a computer ball-and-stick rendering of the methane molecule.



DERIVATIVES OF THE TETRAHEDRON

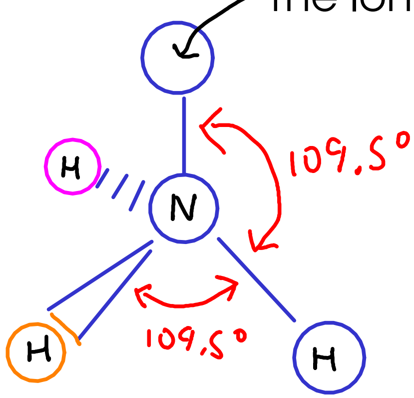
- What if there are lone pairs? The way the shape of a molecule is described depends on the ATOMS in the molecule, even though lone pairs play a role in the positions of the atoms.

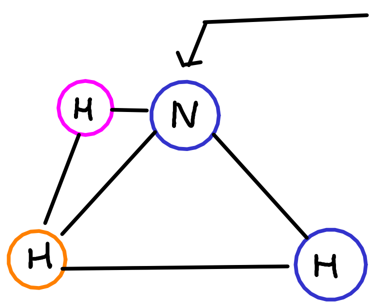


Since there are four "things" around the nitrogen atom, we would expect them to be approximately 109.5 degrees apart (in other words, TETRAHEDRAL). BUT ... only three of these things are atoms.

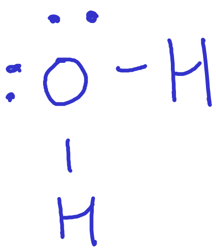
The atoms are arranged in a PYRAMID shape, so we call this molecule PYRAMIDAL!

The lone pair takes one position in the tetrahedron





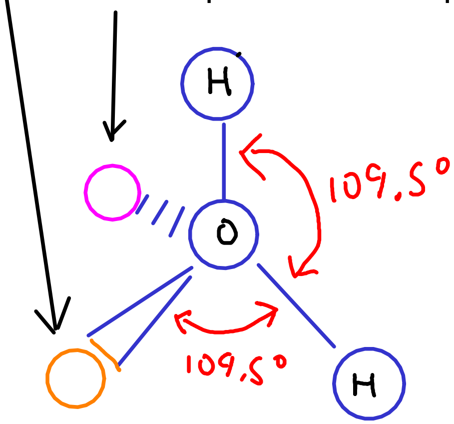
By just looking at the atoms, you can see the pyramid - with the central nitrogen atom as the top and the hydrogen atoms forming the base of the pyramid.

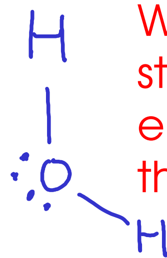


Since there are four "things" around the oxygen atom, we would expect them to be approximately 109.5 degrees apart (in other words, TETRAHEDRAL). BUT... only two of these things are atoms.

The atoms are all in a single plane, but they are not lined up in a straight line. We call this shape "BENT".

Lone pairs take up two positions in the tetrahedron

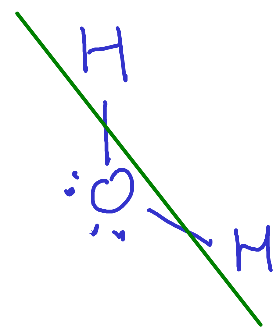




We sometimes draw the Lewis structure of water this way to emphasize the "bent" nature of the molecule!

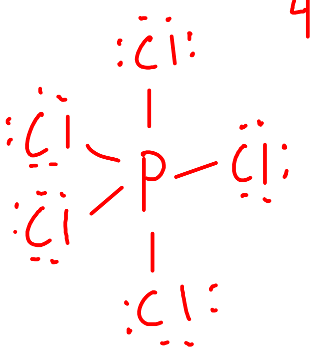
* These atoms are in the same plane, like carbon dioxide. But they are not arranged linearly!

Notice that this molecule has two "sides", one with the oxygen atom and one with hydrogen atoms.

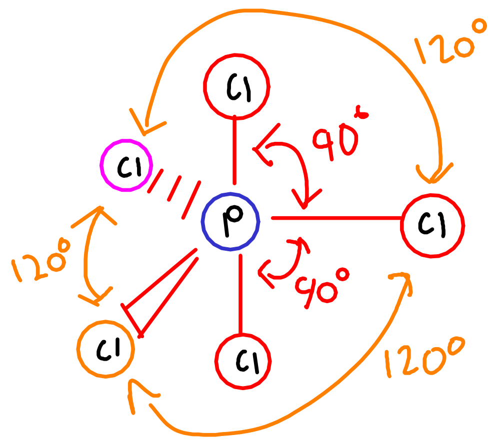


SHAPES OF EXPANDED VALENCE MOLECULES

PCl_5: P : 5
Cl : 7 × 5
40

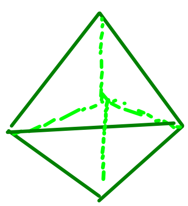


There are five atoms bonded to the central phosphorus atom, and they will attempt to get as far apart as possible from one another!



The top and bottom atoms are 90 degrees apart from the atoms around the center.

The atoms around the center are 120 degrees apart from each other.

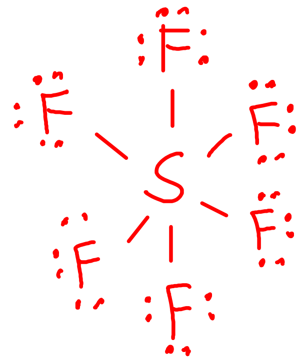


There are acually two DIFFERENT bond angles in this structure. It's called TRIGONAL BIPYRAMIDAL.

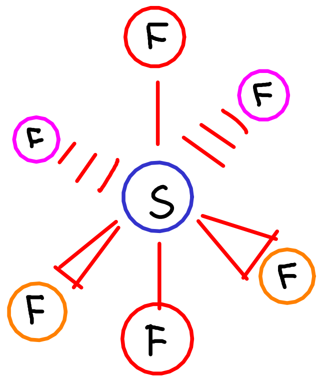
There are several derivatives of the trigonal bipyramidal shape (like the tetrahedral shape) - depending on how many things around the central atom are atoms!

SF_6:

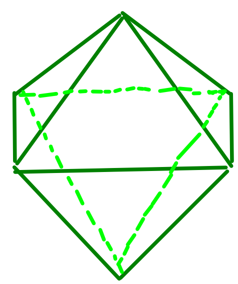
S: 6
F: 7x6
48



There are six atoms bonded to the central sulfur atom, and they will attempt to get as far apart as possible from one another!



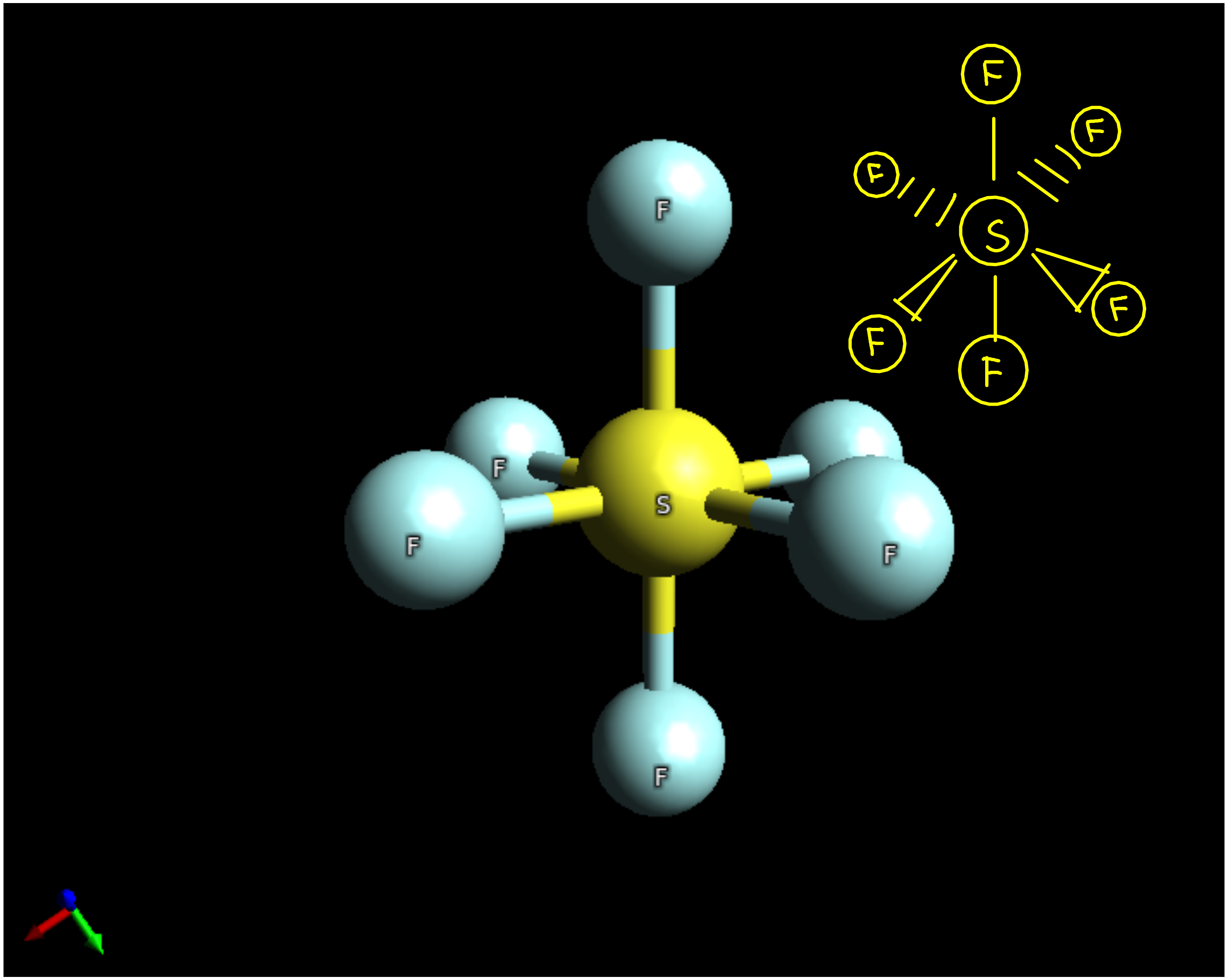
All bond angles in this arrangement are 90 degrees!



This shape is called OCTAHEDRAL, since it has eight sides.

Like the tetrahedral and trigonal bipyramidal arrangements, there are several derivatives of the octahedron - depending on how many of the six things around the center are atoms!

Here's a ball-and-stick rendering of the sulfur hexafluoride molecule:



POLARITY

- When atoms share electrons, the electrons might not be EVENLY shared. Shared electrons may spend more time around one atomic nucleus than the other.

- When electrons are shared UNEVENLY, this results in a POLAR BOND.

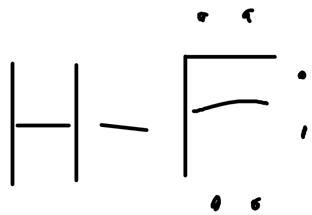
... but how can we tell whether or not a bond will be POLAR? Use ELECTRONEGATIVITY! Usually no actual calculation is required - trends are often good enough to see whether a bond is polar.

REMINDER: ELECTRONEGATIVITY
-A number that describes how tightly an atom will hold shared electrons

- A bond where there is a LARGE electronegativity difference between atoms will be either POLAR or (for very large differences) IONIC!

- A bond with little or no electronegativity difference between atoms will be NONPOLAR

ELECTRONEGATIVITY EXAMPLE

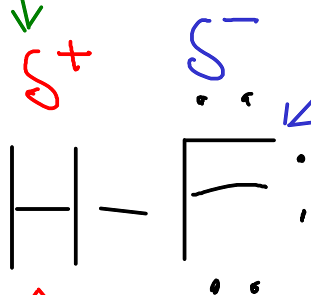


2.1 4.0]– Electronegativity values

Difference = 1.9

Fluorine is much more electronegative than hydrogen. This is a POLAR BOND, and the shared electrons will be held more closely to FLUORINE!

δ means "small" here



This end of the molecule will have a slight NEGATIVE charge, since the shared electrons are closer to FLUORINE!

This end of the molecule will have a slight POSITIVE charge, since the shared electrons are pulled away from HYDROGEN!

POLARITY OF MOLECULES

So what can a molecule's LEWIS STRUCTURE, SHAPE, and the POLARITY of its bonds tell us?

... the POLARITY of the overall molecule, which will tell us (among other things) what a given molecule will mix with or dissolve in!

POLAR MOLECULES

- Will dissolve in or dissolve other polar molecules
- Will dissolve some ionic compounds
- Will NOT easily dissolve nonpolar molecules

Example: WATER

NONPOLAR MOLECULES

- Will dissolve in or dissolve other nonpolar molecules
- Will NOT easily dissolve polar molecules or ionic compounds

Example: OILS

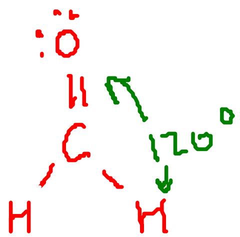
For a molecule to be polar, it must ...

1. Have polar bonds! (Any molecule that contains no polar bonds must be nonpolar!)
2. Have polar bonds arranged in such a way that they don't balance each other out! (This is why you need to know the structure and shape of the molecule)

Examples:

H_2CO

H: 1x2
C: 4
O: 6 | 12



Shape? This molecule is TRIGONAL PLANAR. There are THREE THINGS around the central carbon: =O, -H, and -H

Polar? (1) C=O is polar, C-H isn't. We have at least one polar bond!
(2) Trigonal planar shape. Oxygen has its own "side" of the triangle (which will be negative), and there's nothing but carbon and hydrogen on the other. This is a POLAR molecule.

CO_2

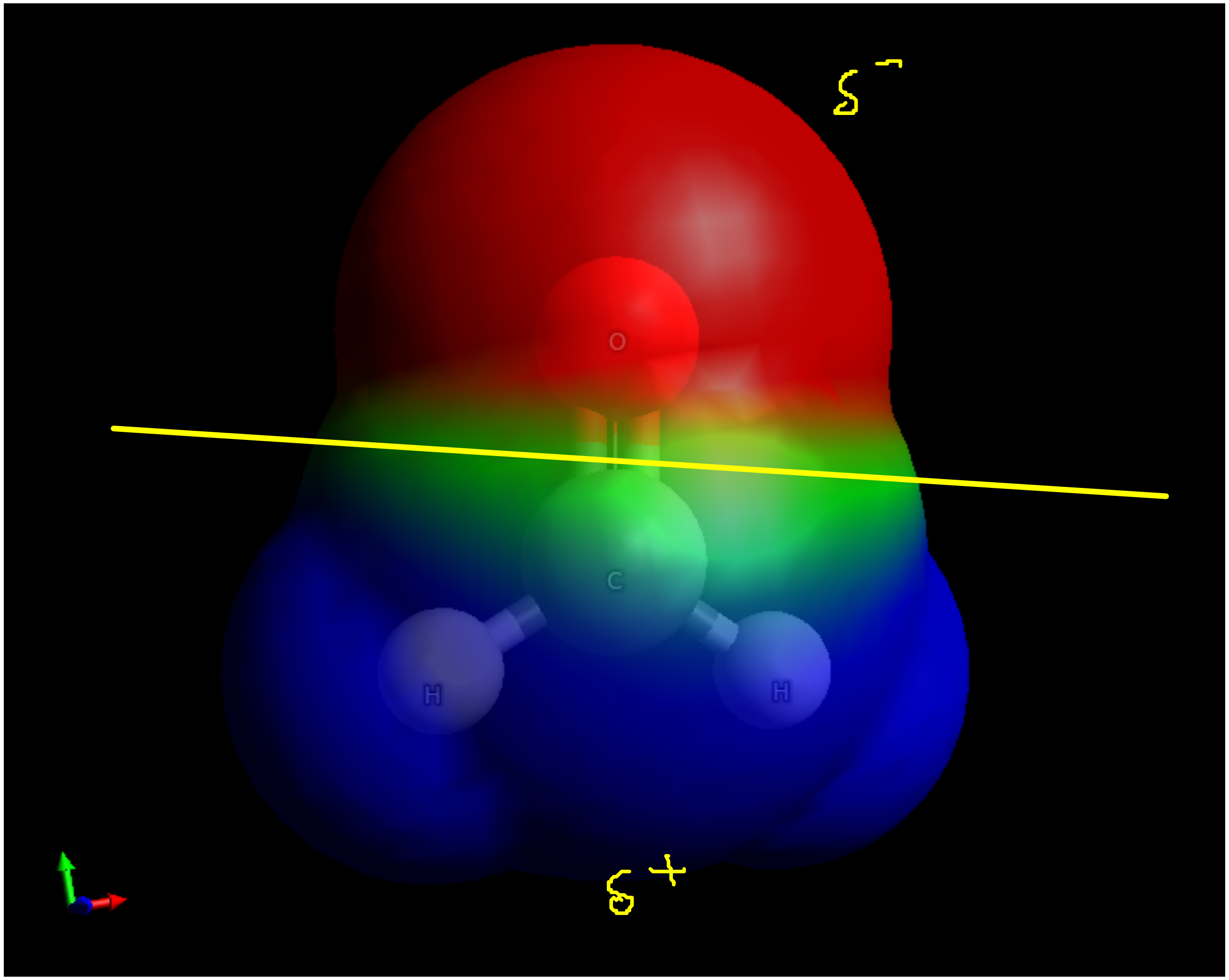
C: 4
O: 6x2
16

Shape? LINEAR. There are only two things around carbon, and they will be 180 degrees apart.

Polar? (1) C=O bonds are polar.
2) Linear shape. The two oxygen atoms directly oppose one another, so there's no positive and negative sides to the molecule. This one is NONPOLAR.

H_2CO

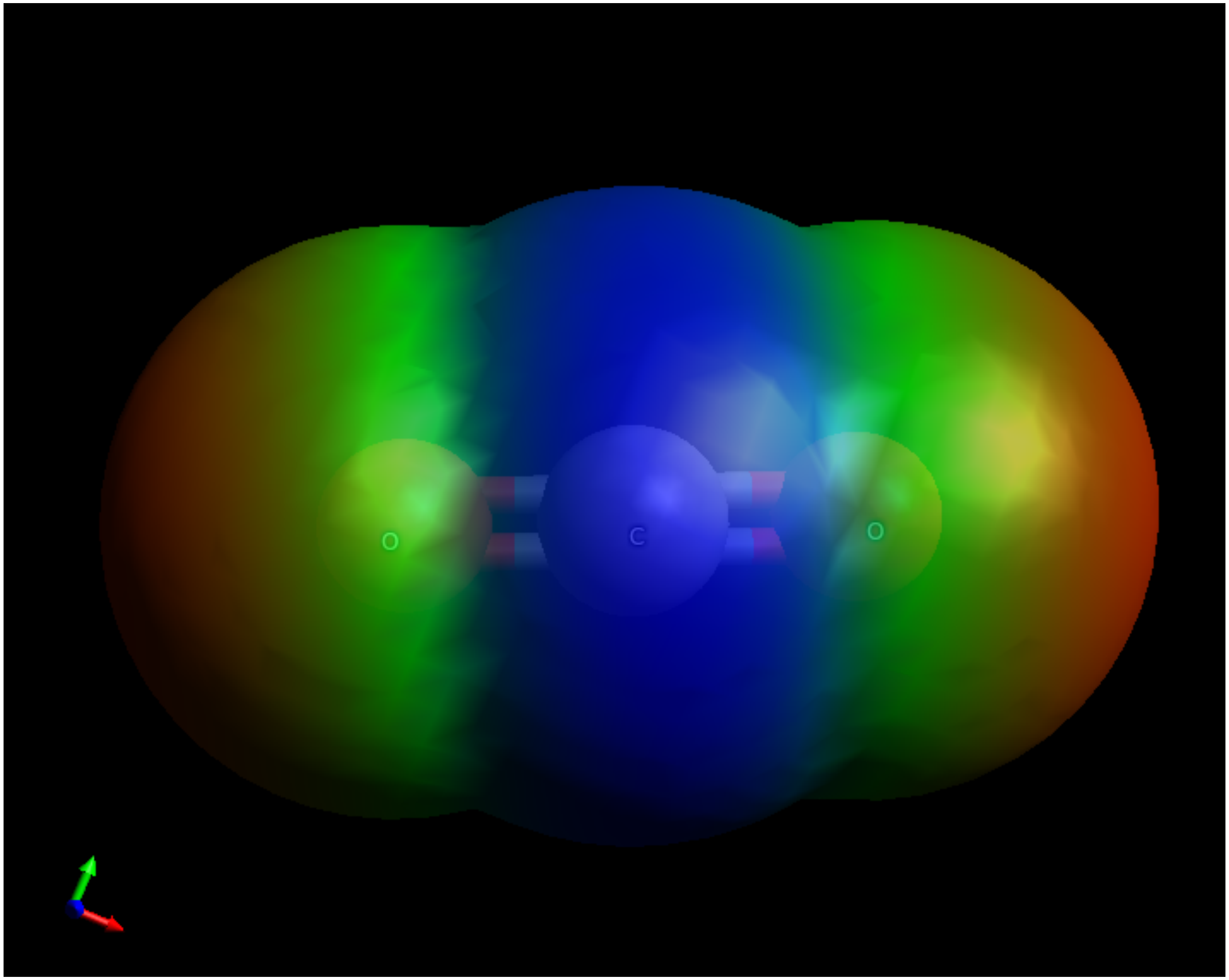
This ball-and-stick model shows electrostatic potential - red for more negative and blue for more positive



oxygen "side", slightly negative

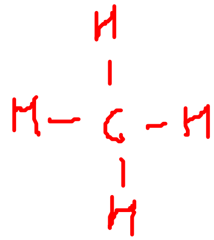
hydrogen "side", slightly positive

CO_2



This molecule is NONPOLAR. No positive "side" or negative "side" methane, CH_4

C: 4
H: 4x1
8

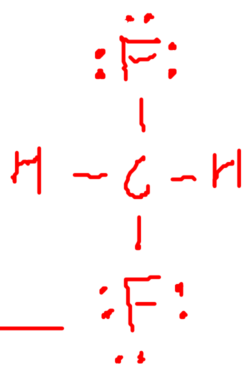


Shape? TETRAHEDRAL. Four atoms around the central carbon atom, all 109.5 degrees apart.

Polar? 1) C-H bonds are nonpolar. Molecule is NONPOLAR.

CH_2F_2

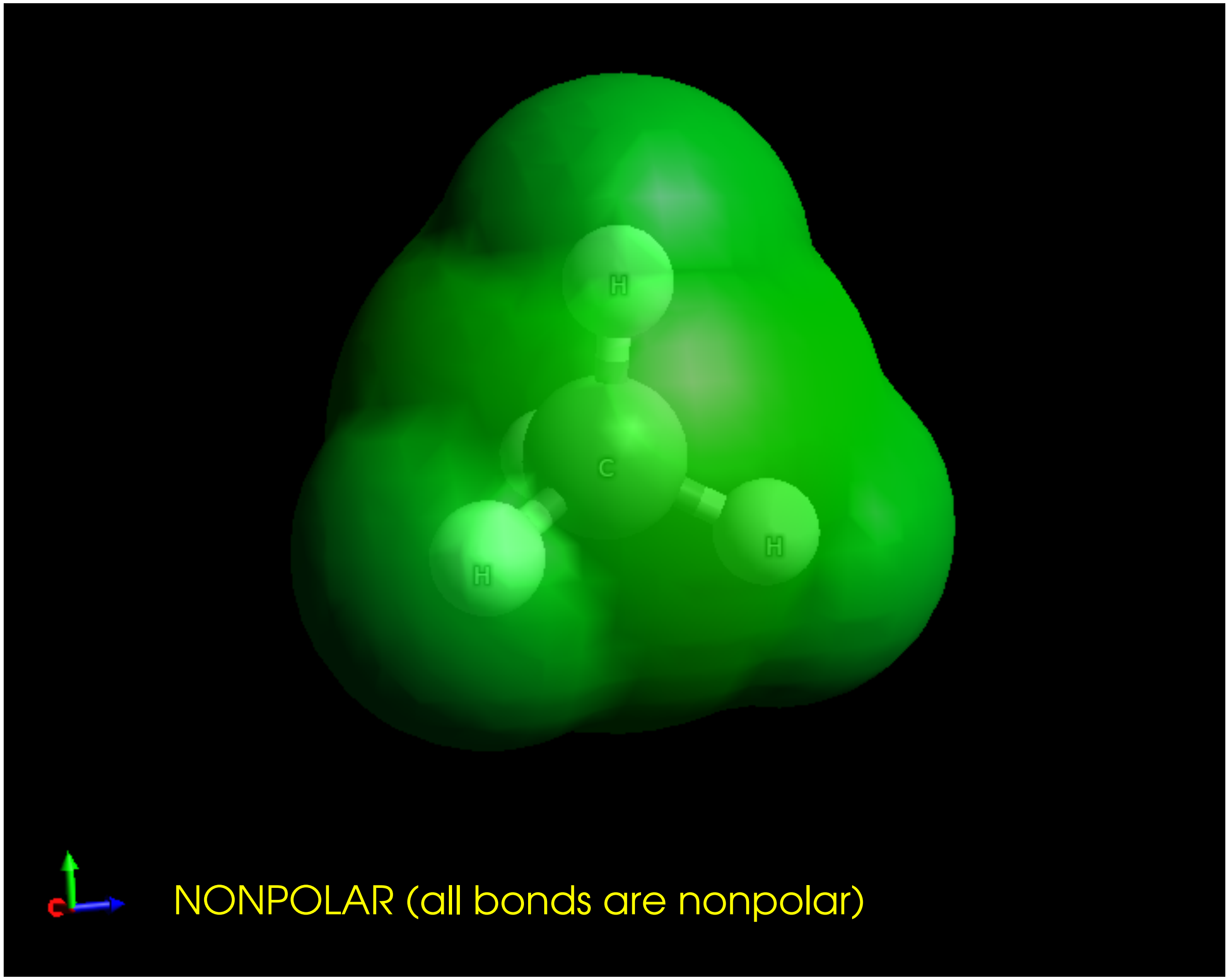
C : 4
H : 2x1 = 2
F : 2x7 = 14
20



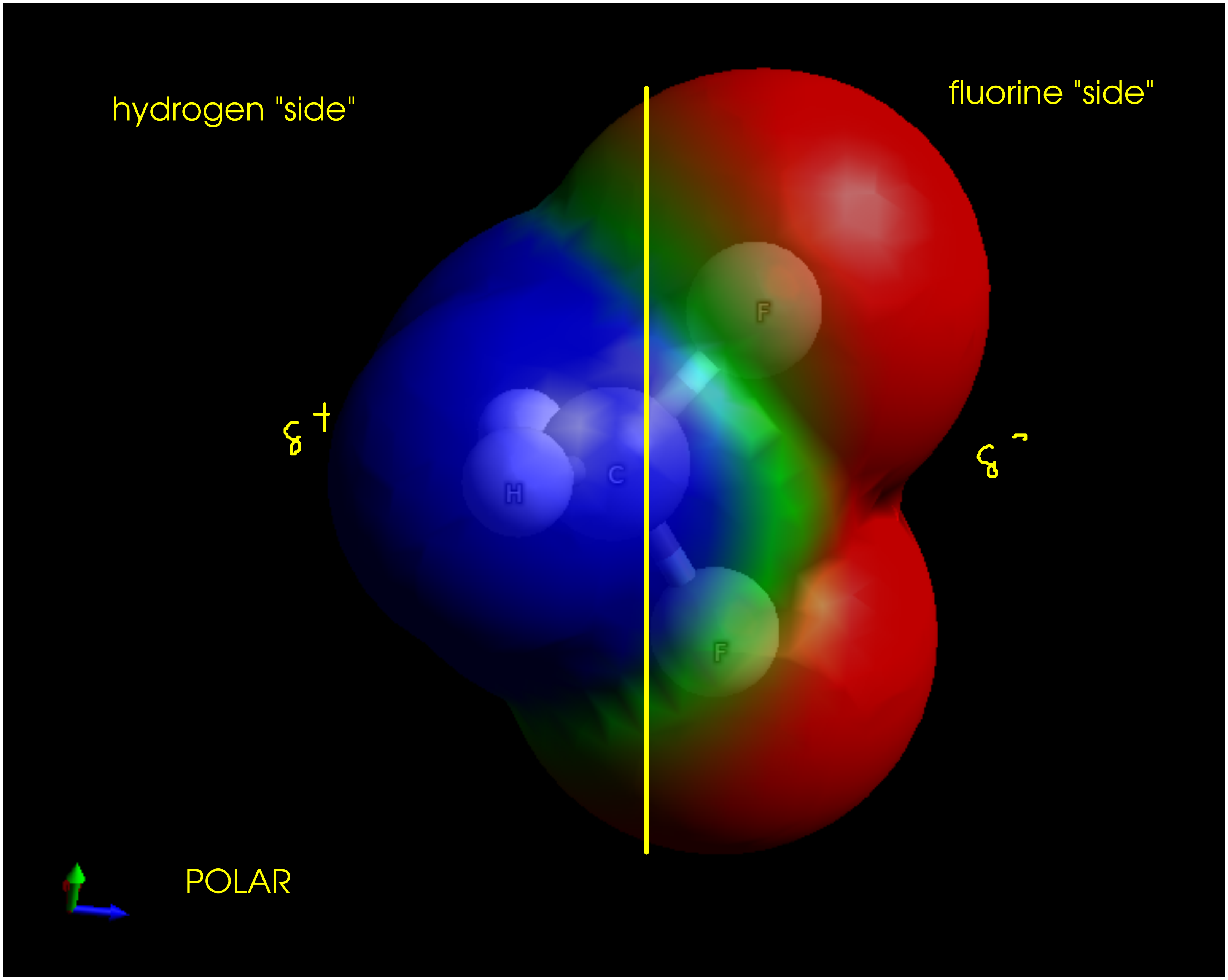
Shape? TETRAHEDRAL. Four atoms around the central carbon atom, just like methane.

Polar? 1) C-H bonds are nonpolar, but C-F bonds are POLAR. We have some polar bonds!
2) Tetrahedral shape, but not all members of the tetrahedron are the same ... so we have two "sides" - a hydrogen side and a fluorine side, The molecule will be POLAR.

CH_4



CH_2F_2



POLARITY AND MOLECULAR PROPERTIES

- POLAR MOLECULES have
    - higher boilng points and melting points that comparably sized nonpolar molecules.
    - higher solubility in polar solvents like water than nonpolar molecules

"LIKE DISSOLVES LIKE"

- NONPOLAR MOLECULES have
    - lower boilng points and melting points that comparably sized polar molecules.
    - higher solubility in nonpolar solvents like carbon tetrachloride or oils