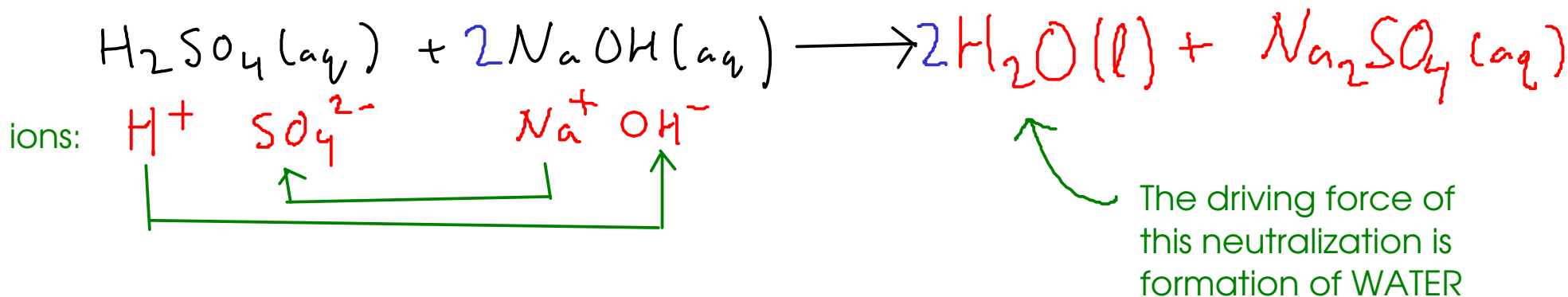
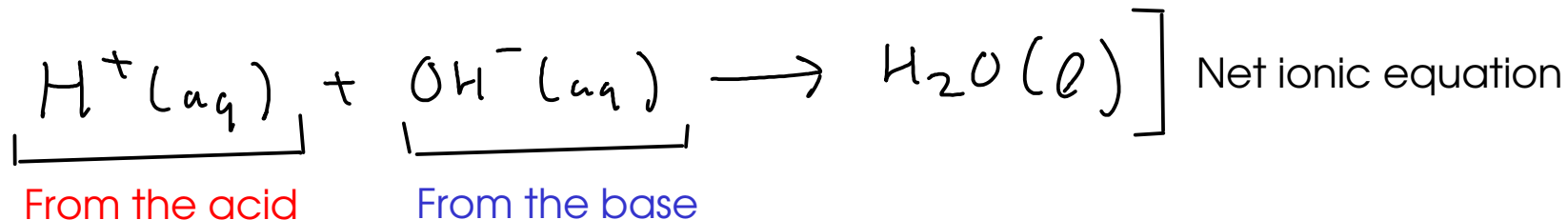


ACID/BASE or NEUTRALIZATION reactions continued

- the driving force of these reactions is the formation of water molecules.



- How can this reaction be detected?

- pH detector (indicator paper, etc.)
- do the products have similar chemical properties to the reactants?
- release of heat!

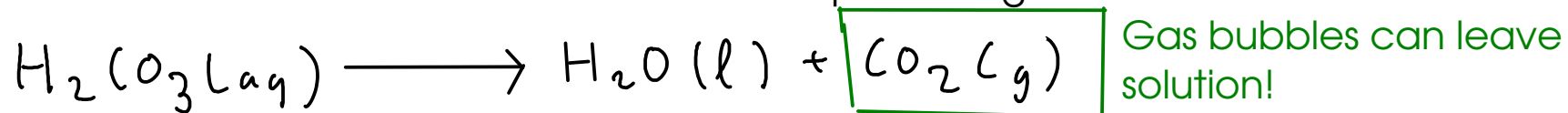
... formation of water is usually accompanied by a release of heat

GAS FORMATION / OTHER MOLECULES

- There are a few other molecules that can be made with exchange-type chemistry.
- Most of these molecules are unstable and can break apart to form gases.

- Formation of a weak acid:

- The formation of ANY weak acid in an exchange-type reaction can be a driving force.
- Some weak acids are unstable and can break apart into gas molecules.

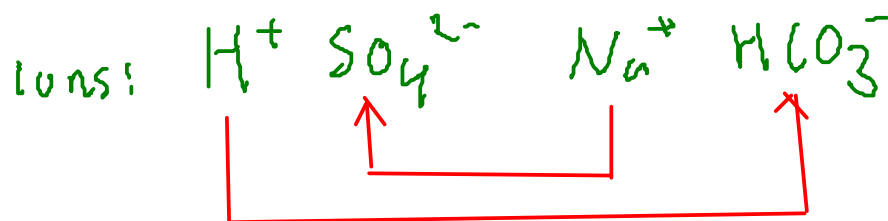
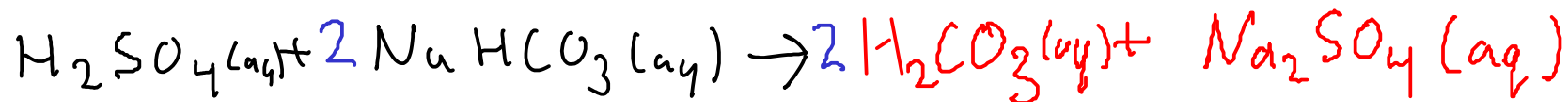


... but how would you form carbonic acid in an exchange-type reaction?

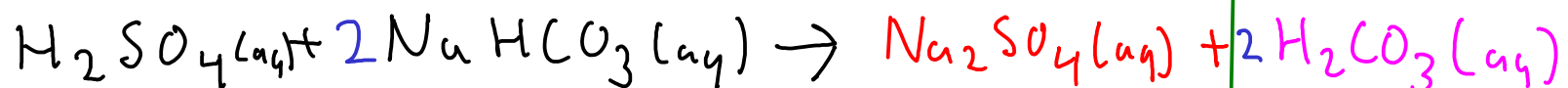
acid + carbonate CO_3^{2-}

OR

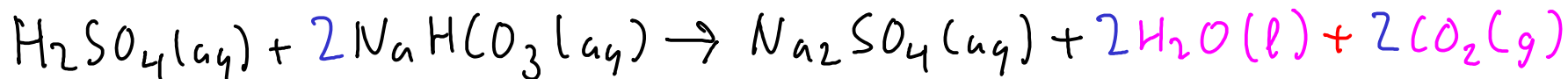
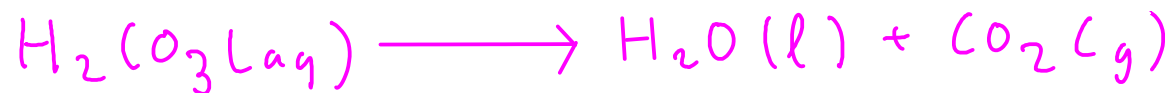
acid + bicarbonate HCO_3^-



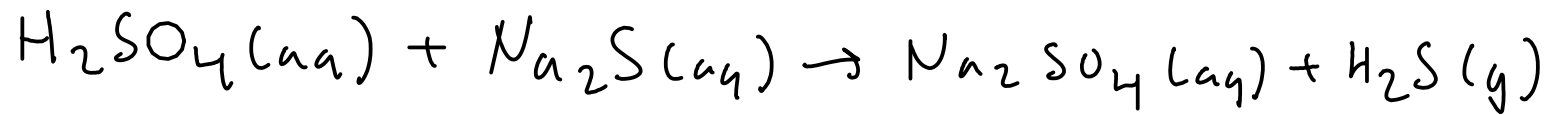
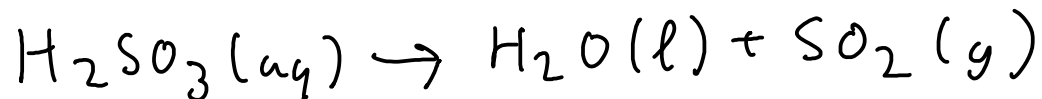
Formation of carbonic acid drives the reaction ... BUT ...



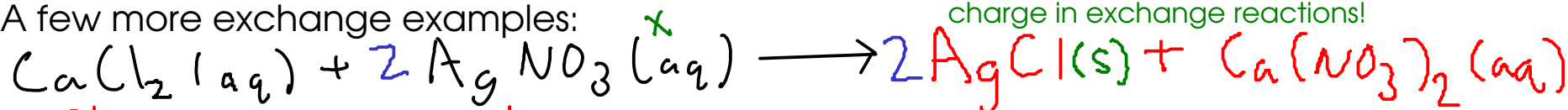
... but when we mix sulfuric acid and sodium bicarbonate, we observe BUBBLES. We need to write an equation that agrees with our observations. We know that carbonic acid decomposes, so we go ahead and put that into our equation.



Other molecules of interest:

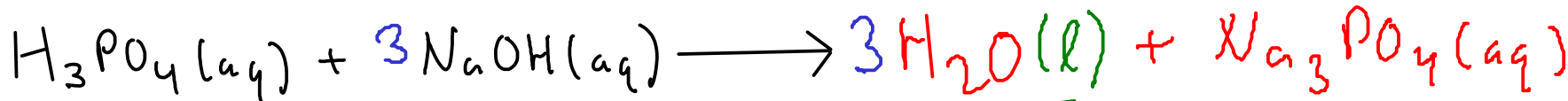


A few more exchange examples:

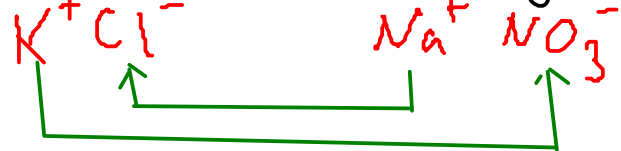
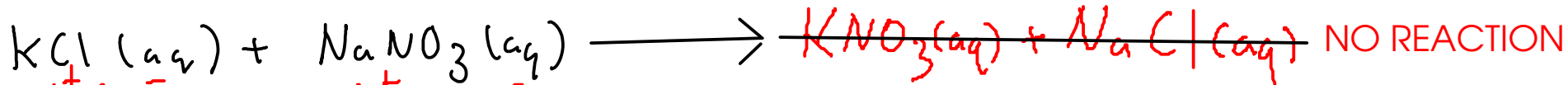


* Transition metals DO NOT change their charge in exchange reactions!

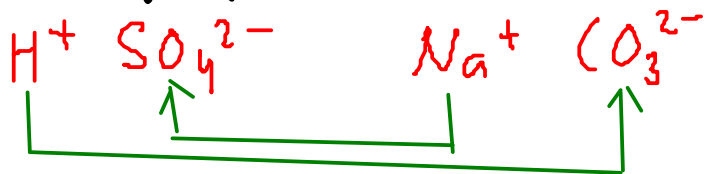
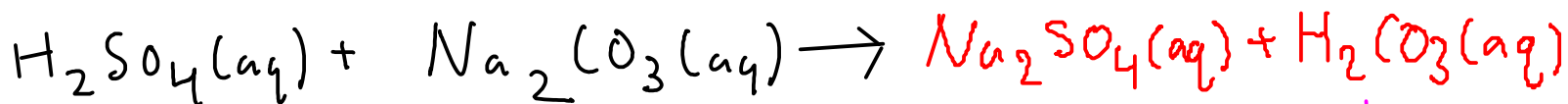
Formation of SOLID AgCl drives this precipitation reaction...



Formation of WATER MOLECULES drives this neutralization reaction ...



Both potential "products" are soluble ionic compounds, meaning that there's no driving force, and NO REACTION. (All the other possible driving forces involve molecules...)



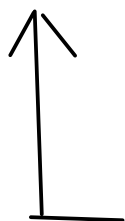
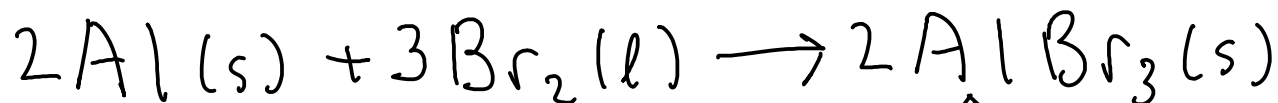
Formation of carbonic acid molecules (and their decomposition into water and carbon dioxide gas) drives this reaction.

125 OXIDATION / REDUCTION CHEMISTRY

- Exchange reactions involve ions pairing up, but the ions themselves are not formed in exchange reactions. Exchanges start with pre-existing ions.

... but the ions have to be produced somehow - through a chemistry that involves the transfer of electrons.

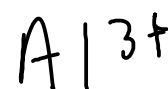
- oxidation / reduction chemistry ("redox" chemistry) involves transfer of electrons and can make ions.



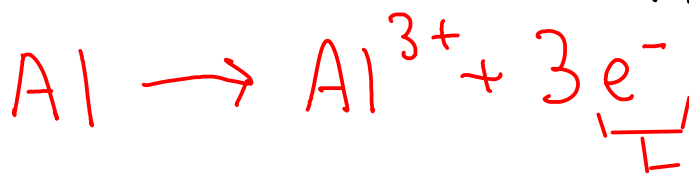
Elemental,
metallic
aluminum.
Uncharged!



Aluminum
cation



These are called
"half-reactions"



electron

oxidation: loss
of electrons



reduction: gain of
electrons