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1) Measure mass of object



2) Partially fill cylinder with liquid, record volume.

3) Put object into cylinder, record new volume

4) Subtract to find volume of object

5) Density = mass object / volume object

Density =
$$\frac{9.78 \text{ g}}{1.6 \text{ mL}}$$

$$= 6.1 \frac{9}{\text{mL}}$$

Converting from one unit to another

We will use the method of dimensional analysis, sometimes called the factor-label method. ... or, the "drag and drop" method!

Dimensional analysis uses conversion factors to change between one unit and another

What's a conversion factor? A simple equality.

Conversion factors in metric

In the metric system, conversion factors between units may always be made from the metric prefixes!

For example, "
$$K_{10}$$
" means 10^{3}
 $K = 10^{3}$
 $K_{m} = 10^{3}$

How do we actually USE a conversion factor?

Convert 15.75 m to ©m
$$Cm = 10^{-2}$$
 m

If $X = 2$, then

$$\frac{X}{2} = \frac{1}{2}$$
IS.75 / EE -2 ... on TI-83

* This fraction equals one, so multiplying by it does not change the VALUE of the number, only its UNITS!

Convert 0.01893 <u>kg</u> to g

$$0.01893 \, \text{kg} \, \text{x} \, \frac{10^3 \, \text{g}}{\text{kg}} = 18.93 \, \text{g}$$

DRAG AND DROP

- Drag the part of the factor that contains the unit you want to get rid of (cancel out) to the BOTTOM.
- Then, drag the other half of the factor to the TOP

Convert 14500 mg to kg
$$mg = 10^{-3}$$
 $(g = 10^{3})$

$$mg = 10^{-3}$$

$$14500 \text{ m/g} \times \frac{10^{-3} \text{g}}{\text{m/g}} \times \frac{\text{kg}}{10^{3} \text{g}} = \boxed{0.0148 \text{ kg}}$$

Convert 0.147 cm² to m²

$$Cm = 10^{2} \text{ Tip: Don't make metric conversion factors out of squared or cubed base units!}$$

$$Cm = 10^{2} \text{ m}$$

When converting squared and cubed units, use each conversion factor two (squared) or three (cubed) times. Remember that

$$(m^2 = (m \times (m + m)))$$

... and it should make sense.

8.45 kg to mg
$$(g - 10^3 g)$$
 $(g - 10^3 g)$

88100 kHz to MHz