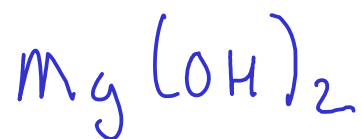


## NAMING IONIC COMPOUNDS

- The name of the compound is based on the name of the ions in the compound
- Cation first, anion second (drop the word "ion")

Examples:



magnesium hydroxide

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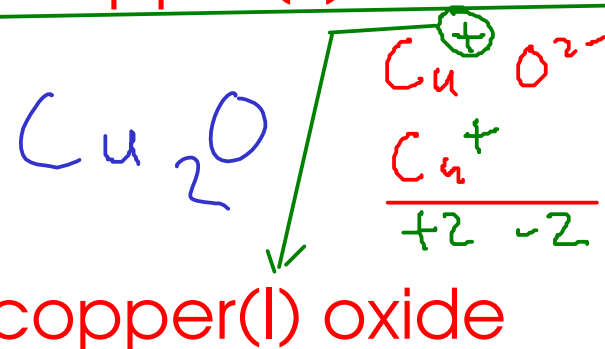
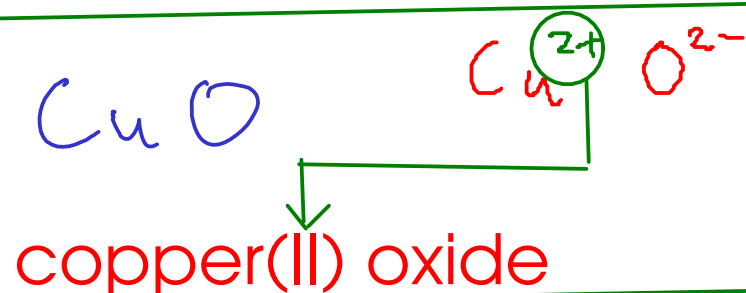
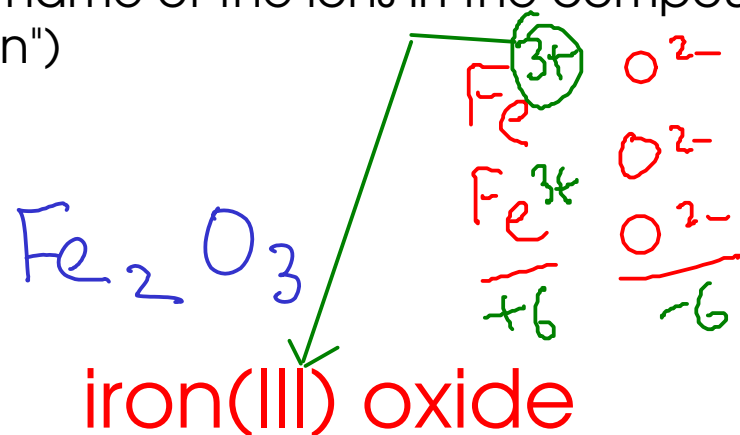
sodium sulfide

---



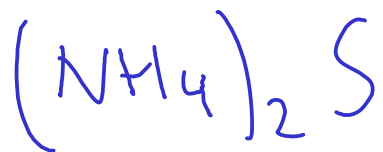
beryllium bromide

---

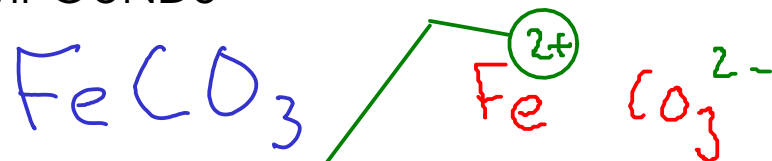


Remember to include the Roman numeral for CHARGE in the name of transition metal compounds!

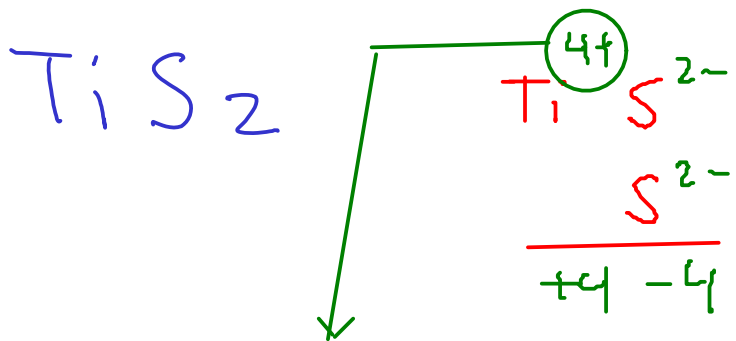
## NAMING IONIC COMPOUNDS



ammonium sulfide



iron(II) carbonate



titanium(IV) sulfide



calcium nitrate



barium phosphate

SPELLING  
MATTERS!



barium phosphide

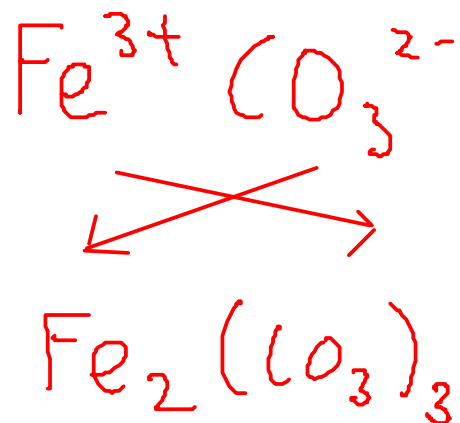
\* p101 - table of polyatomic ions

## DETERMINING THE FORMULA OF AN IONIC COMPOUND FROM THE NAME

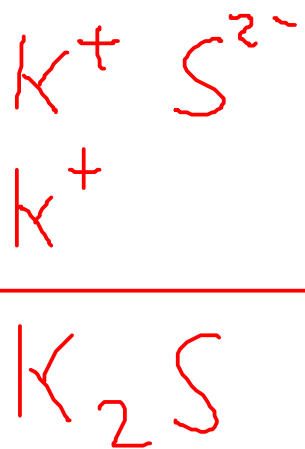
- The name of an ionic compound is made of the names of the CATION and ANION in the compound.
  - To get the FORMULA, you must figure out the SMALLEST RATIO of cation to anion that makes the charges balance out
- 

Examples:

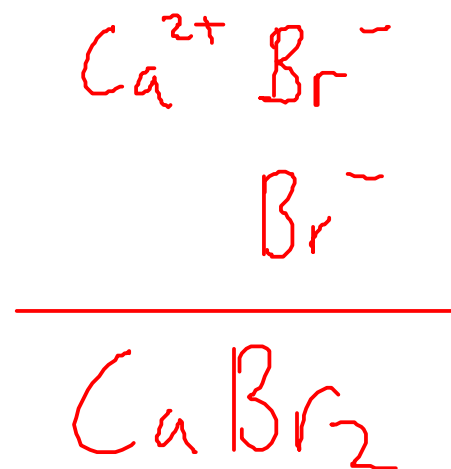
iron(III) carbonate



potassium sulfide

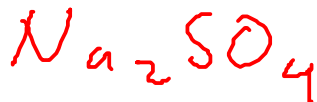


calcium bromide



## DETERMINING IONIC FORMULAS

sodium sulfate



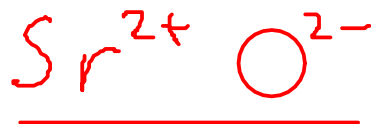
tin(II) phosphate



barium hydroxide



strontium oxide



chromium(III) nitrate



titanium(IV) chloride



chromium(III) nitride



titanium(IV) oxide



Remember to include parenthesis when indicating more than one hydroxide, cyanide, or hypochlorite ion!