

A DOT STRUCTURE FOR A POLYATOMIC ION

① Count valence electrons

② Pick central atom and draw skeletal structure

- central atom is usually the one that needs to gain the most electrons!
- skeletal structure has all atoms connected to center with single bonds

③ Distribute remaining valence electrons around structure, outer atoms first. Follow octet rule until you run out of electrons.

④ Check octet rule - each atom should have a share in 8 electrons (H gets 2). if not, make double or triple bonds.



$$\text{N} : 1 \times 5$$

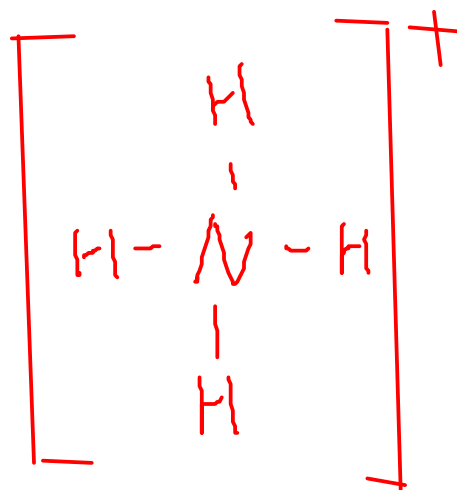
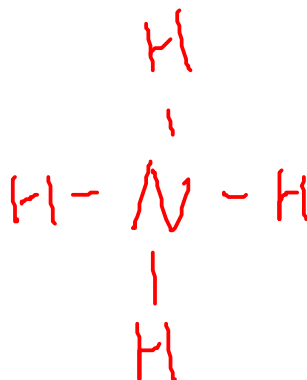
$$\text{H} : 4 \times 1$$

9 electrons

An ODD number of electrons?

- | To get a +1 charge, must LOSE an electron!

8 electrons



To indicate the CHARGE of this ion, draw brackets around the structure and indicate the charge in the upper right corner ...