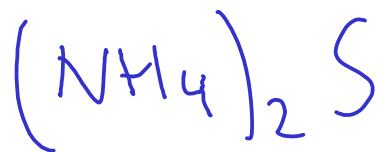
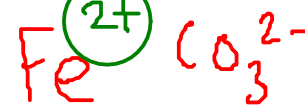


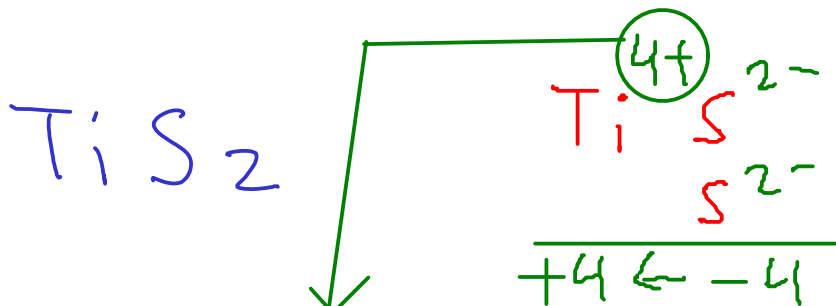
NAMING IONIC COMPOUNDS



ammonium sulfide



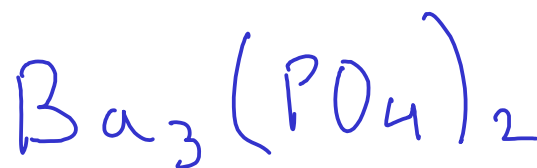
iron(II) carbonate



titanium(IV) sulfide



calcium nitrate



barium phosphate



barium phosphide

SPELLING
MATTERS!

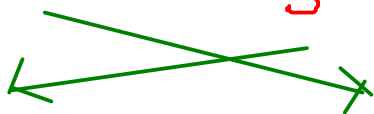
* p130 - table of polyatomic ions

DETERMINING THE FORMULA OF AN IONIC COMPOUND FROM THE NAME

- The name of an ionic compound is made of the names of the CATION and ANION in the compound.
 - To get the FORMULA, you must figure out the SMALLEST RATIO of cation to anion that makes the charges balance out
-

Examples:

iron(III) carbonate



potassium sulfide

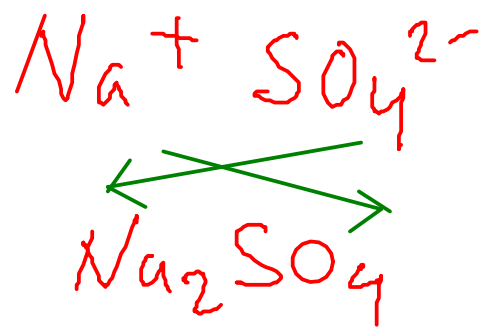


calcium bromide

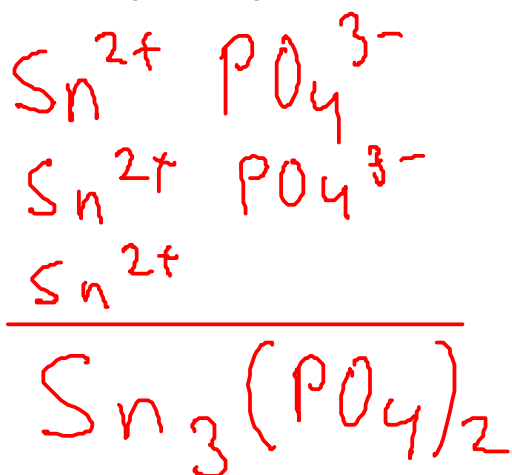


DETERMINING IONIC FORMULAS

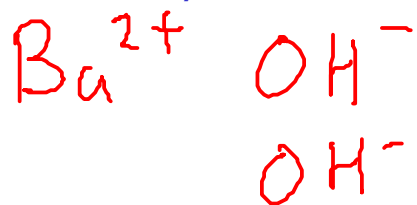
sodium sulfate



tin(II) phosphate



barium hydroxide



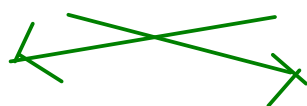
strontium oxide



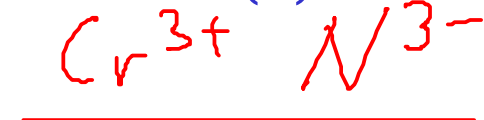
chromium(III) nitrate



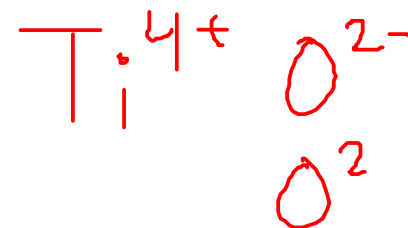
titanium(IV) chloride



chromium(III) nitride



titanium(IV) oxide



Remember to use parenthesis when you have multiple HYDROXIDE, CHLORITE, or CYANIDE ions!