$$2\,NH_4NO_3(s) \longrightarrow \underline{2}\,N_2(g) + \underline{\ }O_2(g) + \underline{4}\,H_2O(g)$$

At 300°C, ammonium nitrate violently decomposes to produce nitrogen gas, oxygen gas, and water vapor. What is the total volume of gas that would be produced at 1.00 atm by the decomposition of 15.0 grams of ammonium nitrate?

To simplify this calculation, calculate the TOTAQL MOLES OF GAS instead of the individual moles of nitrogen gas, oxygen gas, and water vapor.

1- Convert 15.0 grams ammonium nitrate to moles. Use FORMULA WEIGHT.
2 - Convert moles ammnium nitrate to TOTAL MOLES OF GAS using CHEMICAL EQUATION
3 - Convert TOTAL MOLES OF GAS to volume using IDEAL GAS EQUATION

$80.052\text{ g } NH_4NO_3 = \text{mol } NH_4NO_3$ | $2\text{ mol } NH_4NO_3 = 7\text{ mol gas}$ $(2+1+4=7)$

$$15.0\text{ g } NH_4NO_3 \times \underbrace{\frac{\text{mol } NH_4NO_3}{80.052\text{ g } NH_4NO_3}}_{①} \times \underbrace{\frac{7\text{ mol gas}}{2\text{ mol } NH_4NO_3}}_{②} = 0.6558237146\text{ mol gas}$$

③ $PV = nRT$

$V = \dfrac{nRT}{P}$

$n = 0.6558237146\text{ mol gas}$ $\quad R = 0.08206\ \dfrac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$

$T = 300.°C = 573\text{ K}$ $\quad P = 1.00\text{ atm}$

$$V = \frac{(0.6558237146\text{ mol gas})\left(0.08206\ \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}\right)(573\text{ K})}{(1.00\text{ atm})} = \boxed{\begin{array}{c}30.8\text{ L}\\ \text{at } 300°C,\\ 1.00\text{ atm}\end{array}}$$

REAL GASES

- The empirical gas laws (including the ideal gas equation) do not always apply.

  - The gas laws don't apply in situations where the assumptions made by kinetic theory are not valid.

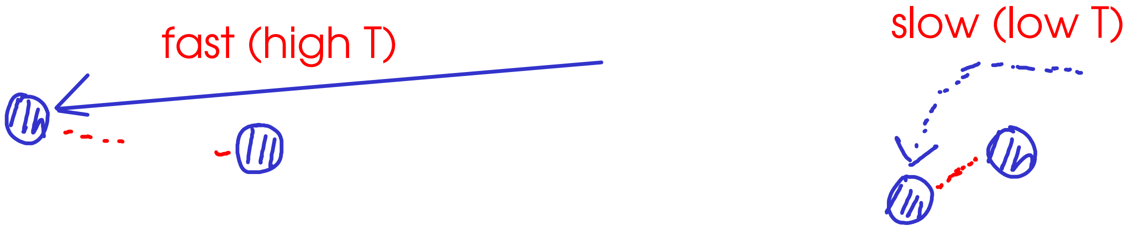
    - When would it be FALSE that the space between gas molecules is much larger than the molecules themselves?

      - at high pressure, molecules would be much closer together!

    - When would it be FALSE that attractive and repulsive forces would be negligible?

      - at high pressure, attractions and repulsions should be stronger!

      - at low temperature, attractions and repulsions have a more significant affect on the paths of molecules



-The gas laws are highly inaccurate near the point where a gas changes to liquid!

- In general, the lower the pressure and the higher the temperature, the more IDEAL a gas behaves.

van der Waals equation

- an attempt to modify PV = nRT to account for several facts.
    - gas molecules actually have SIZE (they take up space)
    - attractive and repulsive forces

$$PV = nRT$$

Ideal gas equation

$$\left(P + \underbrace{\frac{n^2a}{V^2}}\right)\left(V - \underbrace{nb}\right) = nRT$$

van der Waals equation

attempts to account for molecular size

attempts to account for attractive / repulsive forces

* "a" and "b" are experimentally determined parameters that are different for each gas. p 208

He: $a = 0,0346$, $b = 0,0238$ tiny, no special attractive forces

H_2O: $a = 5.537$, $b = 0,03049$ small, but strong attractions between moleculres

CH_3CH_2OH: $a = 12.56$ $b = 0,08710$ larger, and strong attractions between molecules

$25\bar{0}0$ L of chlorine gas at 25.0 C and 1.00 atm are used to make hydrochloric acid. How many kilograms of hydrochloric acid could be produced if all the chlorine reacts?

$$H_2 + Cl_2 \rightarrow 2\,HCl$$

1 - Convert volume of chlorine gas to moles. Use IDEAL GAS EQUATION.
2 - Convert moles chlorine gas to moles HCl. Use CHEMICAL EQUATION/
3 - Convert moles HCl to mass. Use FORMULA WEIGHT. (And g -> kg conversion)

① $PV = nRT$

$n = \frac{PV}{RT}$

$P = 1.00\ \text{atm}$

$V = 25\bar{0}0\ \text{L}$

$R = 0.08206\ \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}$

$T = 25.0^\circ\text{C} = 298.2\ \text{K}$

$$n_{Cl_2} = \frac{(1.00\ \text{atm})(25\bar{0}0\ \text{L})}{\left(0.08206\ \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}}\right)(298.2\ \text{K})} = 102.164983\ \text{mol}\ Cl_2$$

mol Cl_2 = 2 mol HCl | 36.458 g HCl = mol HCl | kg = 10^3 g

$$102.164983\ \text{mol}\ Cl_2 \times \frac{2\ \text{mol HCl}}{\text{mol}\ Cl_2} \times \frac{36.458\ \text{g HCl}}{\text{mol HCl}} \times \frac{\text{kg}}{10^3\ \text{g}} = \boxed{7.45\ \text{kg HCl}}$$

$$2\ HCl + Na_2CO_3 \rightarrow CO_2 + H_2O + 2\ NaCl$$

If 48.90 mL of hydrochloric acid solution react with sodium carbonate to produce 125.0 mL of carbon dioxide gas at 0.950 atm and 290.2 K. What is the molar concentration of the acid?

We need to find M of HCl: $M_{HCl} = \frac{mol\ HCl}{L\ solution} \leftarrow 48.90\ mL = 0.04890\ L$

1 - Convert volume carbn dioxide to moles using IDEAL GAS EQUATION
2 - Convert moles carbon dioxide to moles HCl using CHEMICAL EQUATION
3 - Calculate molarity by dividing moles HCl and volume HCl solution

① $n = \frac{PV}{RT}$

$P = 0.950\ atm \quad V = 125.0\ mL = 0.1250\ L$

$R = 0,08206\ \frac{L\text{-}atm}{mol\text{-}K} \quad T = 290.2\ K$

$$n_{CO_2} = \frac{(0.950\ atm)(0.1250\ L)}{\left(0,08206\ \frac{L\text{-}atm}{mol\text{-}K}\right)(290.2\ K)} = 0,0049866019\ mol\ CO_2$$

$2\ mol\ HCl = mol\ CO_2$

$$0,0049866019\ mol\ CO_2 \times \frac{2\ mol\ HCl}{mol\ CO_2} = 0.0099732038\ mol\ HCl$$

$$M_{HCl} = \frac{mol\ HCl}{L\ solution} = \frac{0.0099732038\ mol\ HCl}{0.04890\ L} = \boxed{0.204\ M\ HCl}$$