MOLAR CONCENTRATION *

- unit: MOLARITY (M): moles of dissolved substance per LITER of solution

dissolved substance

$$M = \text{molarity} = \frac{\text{moles of SOLUTE}}{\text{L SOLUTION}}$$

6.0 <u>M</u> HCl solution: $\frac{6.0 \text{ mol HCl}}{\text{L}}$

If you have 0.250 L (250 mL) of 6.0 M HCl, how many moles of HCl do you have? 6.0 mol HCl = L

$$0.250 \text{ L} \times \frac{6.0 \text{ mol HCl}}{\text{L}} = \boxed{1.5 \text{ mol HCl}}$$

*See SECTIONS 4.7 - 4.10 for more information about MOLARITY and solution calculations (p 154 - 162 - 9th edition) (p 156-164 - 10th edition)

If you need 0.657 moles of hydrochloric acid, how many liters of 0.0555 M HCl do you need to measure out?

$0.0555 \text{ mol HCl} = L$

$$0.657 \text{ mol HCl} \times \frac{L}{0.0555 \text{ mol HCl}} = \boxed{11.8 L}$$

11800 mL

This is too large of a volume for lab-scale work. We probably don't even have this much of the solution available!

What if we used 6.00 M HCl?

$6.00 \text{ mol HCl} = L$

$$0.657 \text{ mol HCl} \times \frac{L}{6.00 \text{ mol HCl}} = \boxed{0.110 L}$$

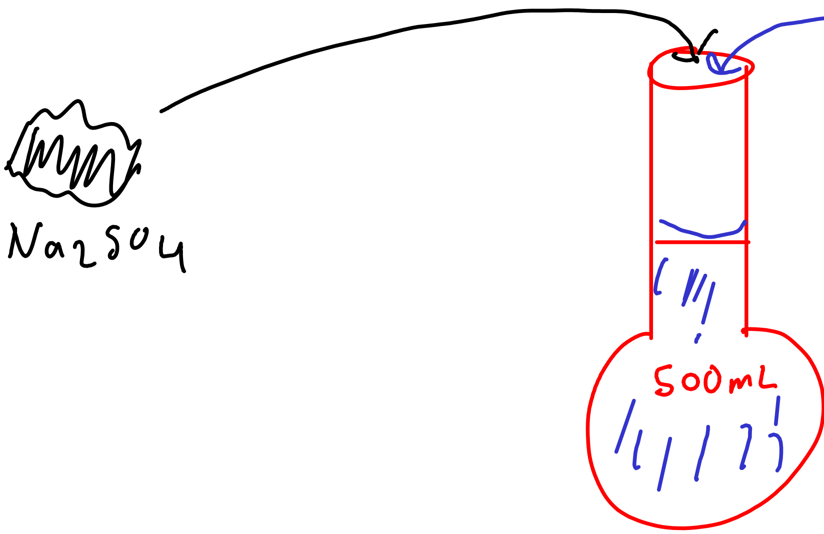
110. mL

A more reasonable lab volume. We'd probably use this (as opposed to the solution above) to get our 0.657 mol HCl

Example: How would we prepare 500. mL of 0.500 M sodium sulfate in water?

Na_2SO_4 : 142.05 g/mol

Dissolve the appropriate amount of sodium sulfate into enough water to make 500. mL of solution.



A VOLUMETRIC FLASK is a flask that is designed to precisely contain a certain volume of liquid.

VOLUMETRIC FLASKS are used to prepare solutions.

volumetric flask

We know that we need 500. mL of solution, and we know that the concentration is supposed to be 0.500 M. We need to figure out how many moles of sodium sulfate should be in that 500 mL solution. If we convert that number of moles to mass, we'll know how much sodium sulfate to weight out!

$0.500 \text{ mol } Na_2SO_4 = L$	$mL = 10^{-3} L$	$142.05 \text{ g } Na_2SO_4 = \text{mol } Na_2SO_4$

$$500. \cancel{mL} \times \frac{10^{-3} \cancel{L}}{\cancel{mL}} \times \frac{0.500 \cancel{\text{mol } Na_2SO_4}}{\cancel{L}} \times \frac{142.05 \text{ g } Na_2SO_4}{\cancel{\text{mol } Na_2SO_4}} = \boxed{35.5 \text{ g } Na_2SO_4}$$

So, to prepare this solution, we would add 35.5 grams of sodium sulfate to a 500 mL volumetric flask and dilute to the mark with distilled water.

More on MOLARITY

To prepare a solution of a given molarity, you generally have two options:

1. Weigh out the appropriate amount of solute, then dilute to the desired volume with solvent (usually water)

"stock solution"

2. Take a previously prepared solution of known concentration and DILUTE it with solvent to form a new solution

- Use DILUTION EQUATION

The dilution equation is easy to derive with simple algebra.

$$M \times V$$

$$\frac{mol}{L} \times L = \text{moles solute}$$

... but when you dilute a solution, the number of moles of solute REMAINS CONSTANT. (After all, you're adding only SOLVENT)

$$M_1 V_1 = M_2 V_2$$

Since the number of moles of solute stays the same, this equality must be true!

before diution

after dilution

$$M_1 V_1 = M_2 V_2$$ ... the "DILUTION EQUATION"

M_1 = molarity of concentrated solution

V_1 = volume of concentrated solution

M_2 = molarity of dilute solution

V_2 = volume of dilute solution (total volume, not volume of added solvent!)

The volumes don't HAVE to be in liters, as long as you use the same volume UNIT for both volumes!

Example: Take the 0.500 M sodium sulfate we discussed in the previous example and dilute it to make 150. mL of 0.333 M solution. How many mL of the original solution will we need to dilute?

$$M_1 V_1 = M_2 V_2$$

$M_1 = 0.500\,M$ | $M_2 = 0.333\,M$

$V_1 = ?$ | $V_2 = 150.\,mL$

$$(0.500\,M) \times V_1 = (0.333\,M)(150.\,mL)$$

$$\boxed{V_1 = 99.9\,mL \text{ of } 0.500\,M\ Na_2SO_4}$$

So, to malke the solution, take 99.9 mL of 0.500 M sodium sulfate and add enough distilled water to make 150. mL of solution.

CHEMICAL CALCULATIONS CONTINUED: REACTIONS

- Chemical reactions proceed on an ATOMIC basis, NOT a mass basis!

- To calculate with chemical reactions (i.e. use chemical equations), we need everything in terms of ATOMS ... which means MOLES of atoms

$$2\,Al(s) + 3\,Br_2(l) \longrightarrow 2\,AlBr_3(s)$$

coefficients are in terms of atoms and molecules!

2 atoms Al = 3 molecules Br_2 = 2 formula units $AlBr_3$

2 mol Al = 3 mol Br_2 = 2 mol $AlBr_3$

- To do chemical calculations, we need to:
  - Relate the amount of substance we know (mass or volume) to a number of moles
  - Relate the moles of one substance to the moles of another using the equation
  - Convert the moles of the new substance to mass or volume as desired

$$2Al(s) + 3Br_2(l) \longrightarrow 2AlBr_3(s)$$

* Given that we have 25.0 g of liquid bromine, how many grams of aluminum would we need to react away all of the bromine? How many grams of aluminum bromide would be produced?

1. Convert grams of bromine to moles: Need formula weight Br_2: $\frac{2 \times 79.90}{159.80}$

$154.80\,g\,Br_2 = 1\,mol\,Br_2$

$$25.0\,g\,Br_2 \times \frac{1\,mol\,Br_2}{154.80\,g\,Br_2} = 0.15645\,mol\,Br_2$$

2. Use the chemical equation to relate moles of bromine to moles of aluminum

$2\,mol\,Al = 3\,mol\,Br_2$

$$0.15645\,mol\,Br_2 \times \frac{2\,mol\,Al}{3\,mol\,Br_2} = 0.10430\,mol\,Al$$

3. Convert moles aluminum to mass: Need formula weight Al: 26.98

$26.98\,g\,Al = 1\,mol\,Al$

$$0.10430\,mol\,Al \times \frac{26.98\,g\,Al}{1\,mol\,Al} = \boxed{2.81\,g\,Al}$$

You can combine all three steps on one line if you like!

$$25.0\,g\,Br_2 \times \frac{1\,mol\,Br_2}{159.80\,g\,Br_2} \times \frac{2\,mol\,Al}{3\,mol\,Br_2} \times \frac{26.98\,g\,Al}{1\,mol\,Al} = \boxed{2.81\,g\,Al}$$

1 2 3

You can solve the second part of the question using CONSERVATION OF MASS - since there's only a single product and you already know the mass of all reactants.

$$\begin{array}{r} 25.0\,g\,Br_2 \\ +\ 2.81\,g\,Al \\ \hline 27.8\,g\,AlBr_3 \end{array}$$

But ...

...what would you have done to calculate the mass of aluminum bromide IF you had NOT been asked to calculate the mass of aluminum FIRST?

$$25.0\,g\,Br_2 \times \frac{1\,mol\,Br_2}{159.80\,g\,Br_2} \times \frac{2\,mol\,AlBr_3}{3\,mol\,Br_2} \times \frac{266.694\,g\,AlBr_3}{1\,mol\,AlBr_3} = 27.8\,g\,AlBr_3$$

1 convert mass bromine to moles

2 convert moles bromine to moles aluminum bromide

3 convert moles aluminum bromide to mass

Example:

How many milliliters of 6.00M hydrochloric acid is needed to completely react with 25.0 g of sodium carbonate?

$$2HCl(aq) + Na_2CO_3(s) \longrightarrow H_2O(l) + CO_2(g) + 2NaCl(aq)$$

1 - Convert 25.0 g sodium carbonate to moles. Use formula weight.
2 - Convert moles sodium carbonate to moles HCl. Use chemical equation.
3 - Convert moles HCl to volume. Use MOLAR CONCENTRATION (6.00 mol HCl = L)

1 Na_2CO_3: Na : 2 x 22.99
C : 1 x 12.01
O : 3 x 16.00

105.99 g Na_2CO_3 = mol Na_2CO_3

$$25.0\text{ g } Na_2CO_3 \times \frac{\text{mol } Na_2CO_3}{105.99\text{ g } Na_2CO_3} = 0.2358713086\text{ mol } Na_2CO_3$$

2 2 mol HCl = mol Na_2CO_3

$$0.2358713086\text{ mol } Na_2CO_3 \times \frac{2\text{ mol HCl}}{\text{mol } Na_2CO_3} = 0.4717426172\text{ mol HCl}$$

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1 - Convert 25.0 g sodium carbonate to moles. Use formula weight.
2 - Convert moles sodium carbonate to moles HCl. Use chemical equation.
3 - Convert moles HCl to volume. Use MOLAR CONCENTRATION (6.00 mol HCl = L)

3 $6.00 \text{ mol HCl} = L$ | $mL = 10^{-3} L$

$$0.4717426172 \text{ mol HCl} \times \frac{L}{6.00 \text{ mol HCl}} \times \underbrace{\frac{mL}{10^{-3} L}} = \boxed{78.6 \text{ mL of } 6.00M \text{ HCl}}$$

We used this factor to convert L of solution to mL, since the problem specifically asks us for mL!