Water:

ENERGY = MASS $\times$ SPECIFIC HEAT $\times$ TEMPERATURE CHANGE

$(28.9°C - 22.2°C)$

ENERGY = $\underline{50.0\ g}$ $\times$ $4.184\ \dfrac{J}{g°C}$ $\times$ $\underline{6.7}$ $°C$

$= 1401.64\ J$

This number has two significant figures, but we'll wait until the final answer to round!

By conservation of energy, this energy gained by the water is also equal to the energy LOST be the zinc!

Zinc:

SPECIFIC HEAT $= \dfrac{ENERGY}{MASS \times TEMPERATURE\ CHANGE}$

The temp change of the metal is different from the temp change of the water!

$99.7°C$
$- 28.9°C$
$70.8°C$

SPECIFIC HEAT $= \dfrac{1401.64\ J}{50.2871\ g \times \underline{70.8}\ °C}$

$= \boxed{0.39 \quad J/g°C}$

$0.393682947 52\ ^J/_{g°C}$