

If 50.3 kJ of heat was released when 5.48 g of formic acid are burned at constant pressure, then what is the enthalpy change of this reaction per mole of formic acid?

$$Q = -50.3 \text{ kJ} \quad \Delta H = \frac{Q_{\text{constant pressure}}}{\text{mol formic acid}}$$

Find moles formic acid:

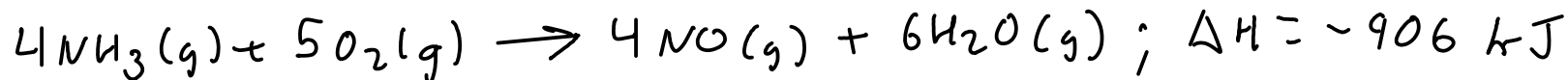
$$5.48 \text{ g HCHO}_2 \times \frac{\text{mol HCHO}_2}{46.026 \text{ g HCHO}_2} = 0.1190631382 \text{ mol HCHO}_2$$

$$\Delta H = \frac{-50.3 \text{ kJ}}{0.1190631382 \text{ mol HCHO}_2} = \boxed{-422 \frac{\text{kJ}}{\text{mol HCHO}_2}}$$

Based on the calculation above, can we complete this thermochemical equation?



We calculated the enthalpy change PER MOLE of formic acid (as is normal in calorimetry experiments), but the thermochemical equation written here is based on TWO MOLES of formic acid. So the enthalpy change for the equation is twice what we calculated in the example above: -844 kJ instead of -422 kJ.



What is the enthalpy change when 150. L of nitrogen monoxide are formed by this reaction at 25.0 C and 1.50 atm pressure?

1 - Convert 150 L NO to moles NO using ideal gas equation.

2 - Convert moles NO to enthalpy change using thermochemical equation.

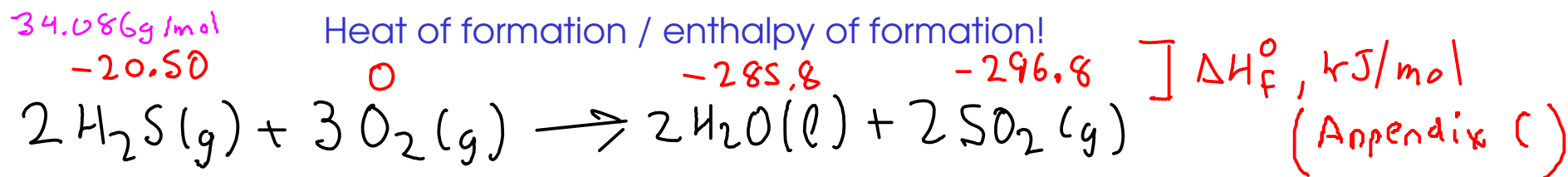
$$PV = nRT \quad \left| \quad P = 1.50 \text{ atm} \quad V = 150. \text{ L} \right.$$

$$n = \frac{PV}{RT} \quad \left| \quad R = 0.08206 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \quad T = 25.0^\circ\text{C} = 298.2 \text{ K} \right.$$

$$\textcircled{1} n_{\text{NO}} = \frac{(1.50 \text{ atm})(150. \text{ L})}{\left(0.08206 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}\right)(298.2 \text{ K})} = 9.194822849 \text{ mol NO}$$

$$4 \text{ mol NO} = -906 \text{ kJ}$$

$$9.194822849 \text{ mol NO} \times \frac{-906 \text{ kJ}}{4 \text{ mol NO}} = \boxed{-2080 \text{ kJ}}$$



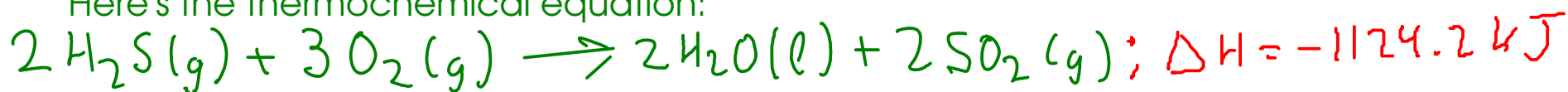
What is the enthalpy change at standard conditions when 25.0 grams of hydrogen sulfide gas is reacted?

- 1 - Calculate the enthalpy change for the reaction as written using HEATS OF FORMATION
- 2 - Convert 25.0 grams of hydrogen sulfide to moles using formula weight.
- 3 - Convert moles hydrogen sulfide to enthalpy change using thermochemical equation.

$$\Delta H = \sum \Delta H_f^\circ, \text{products} - \sum \Delta H_f^\circ, \text{reactants}$$

$$\textcircled{1} = 2(-285.8) + 2(-296.8) - [2(-20.50) + 3(0)] = -1124.2 \text{ kJ}$$

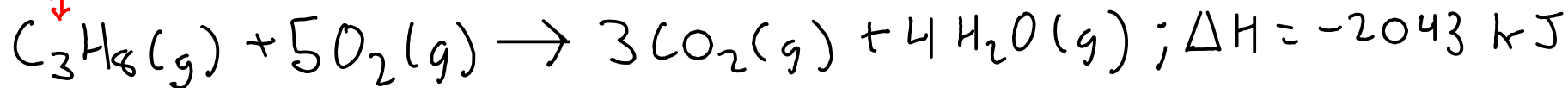
Here's the thermochemical equation:



$$34.086 \text{ g H}_2\text{S} = 1 \text{ mol H}_2\text{S} \quad | \quad 2 \text{ mol H}_2\text{S} = -1124.2 \text{ kJ}$$

$$25.0 \text{ g H}_2\text{S} \times \frac{1 \text{ mol H}_2\text{S}}{34.086 \text{ g H}_2\text{S}} \times \frac{-1124.2 \text{ kJ}}{2 \text{ mol H}_2\text{S}} = \boxed{-412 \text{ kJ}}$$

propane



Calculate the volume of propane gas at 25.0 C and 1.08 atm required to provide 565 kJ of heat using the reaction above.

- 1 - Convert energy requirement to moles propane using thermochemical equation.
- 2 - Convert moles propane to volume using ideal gas equation.

$$\text{mol C}_3\text{H}_8 = -2043 \text{ kJ}$$

Since the reaction is providing the energy, the energy requirement is NEGATIVE from the point of view of the reaction!

$$-565 \text{ kJ} \times \frac{\text{mol C}_3\text{H}_8}{-2043 \text{ kJ}} = 0.2765540871 \text{ mol C}_3\text{H}_8 \quad (1)$$

$$PV = nRT \quad \left| \quad n = 0.2765540871 \text{ mol C}_3\text{H}_8 \quad T = 25.0^\circ\text{C} = 298.2 \text{ K}$$

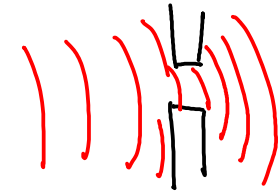
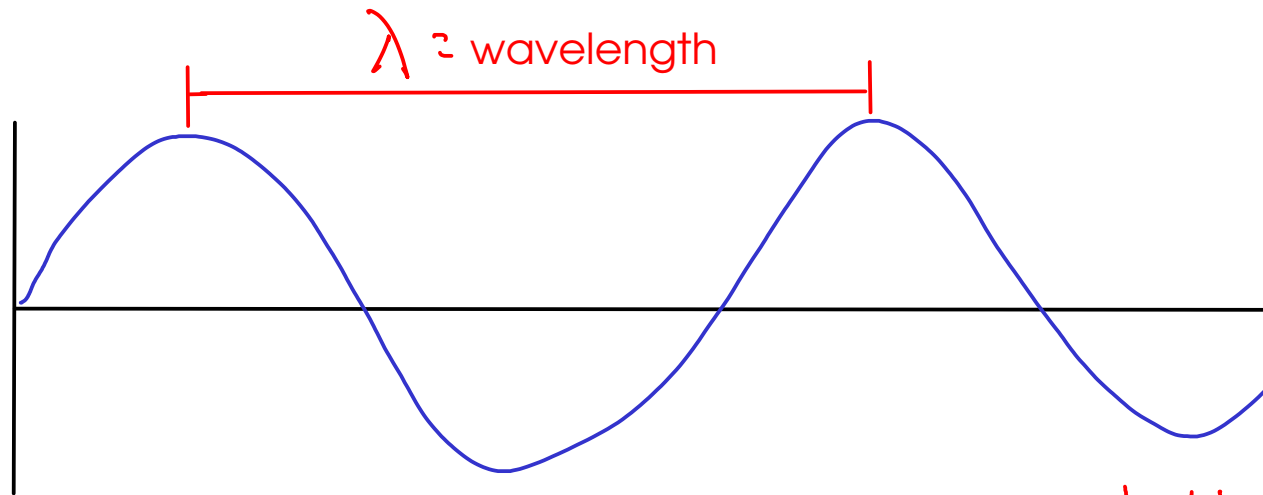
$$V = \frac{nRT}{P} \quad \left| \quad R = 0.08206 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}} \quad P = 1.08 \text{ atm}$$

$$V = \frac{(0.2765540871 \text{ mol C}_3\text{H}_8)(0.08206 \frac{\text{L}\cdot\text{atm}}{\text{mol}\cdot\text{K}})(298.2 \text{ K})}{(1.08 \text{ atm})} =$$

$$= \boxed{6.27 \text{ L propane @ } 25.0^\circ\text{C} \text{ } 1.08 \text{ atm}}$$

END OF CHAPTER 6

LIGHT



Diffraction

$$\text{frequency} = \text{wavelengths} / \text{time} = \nu \quad \text{s}^{-1} : \text{Hertz}, \text{ Hz}$$

- Light has properties of WAVES such as DIFFRACTION (it bends around small obstructions).
- Einstein noted that viewing light as a particle that carried an energy proportional to the FREQUENCY could explain the PHOTOELECTRIC EFFECT!

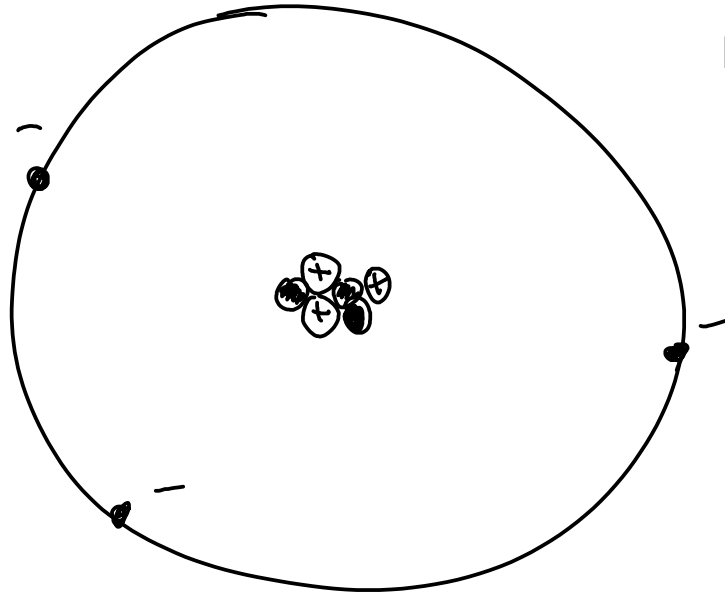
$$E_{\text{photon}} = h \nu$$

Planck's constant: $6.63 \times 10^{-34} \text{ J}\cdot\text{s}$

photon = particle or packet of light

(The photoelectric effect is the emission of electrons from a metal caused by exposure to light. Einstein discovered that if the light were not of the correct FREQUENCY, increasing the INTENSITY of the light would not cause electron emission. He concluded that individual photons must have enough energy to excite an electron - i.e. they must have the appropriate frequency.)

The photoelectric effect and Einstein's ideas about the energy content of light led us to discover a new model for the atom! How? Let's start with the nuclear model:



Nuclear model:

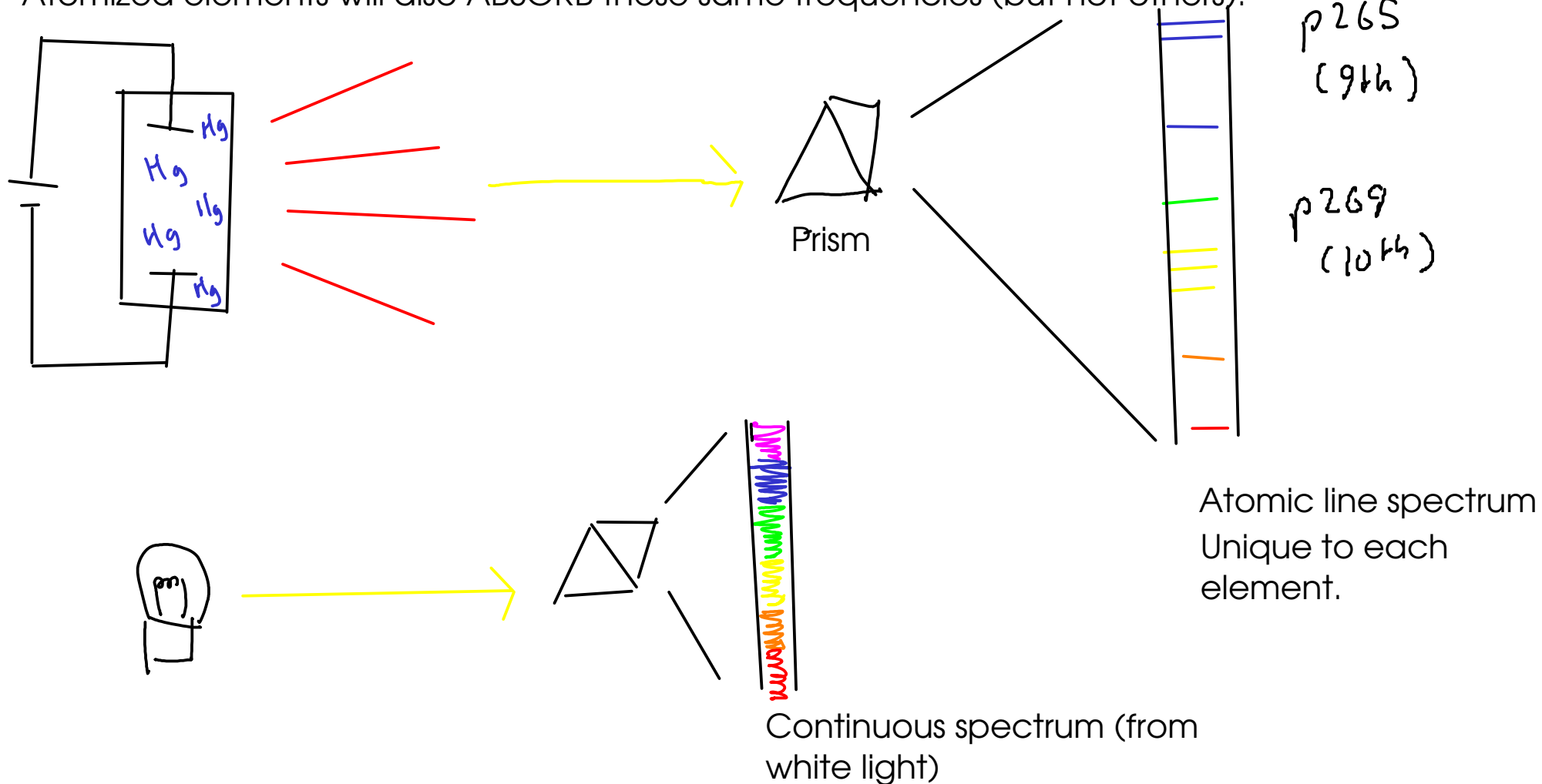
- Protons and neutrons in a dense NUCLEUS at center of atom
- Electrons in a diffuse (mostly empty) ELECTRON CLOUD surrounding NUCLEUS.

... so what's wrong with the nuclear model? Among other things, it doesn't explain ...

ATOMIC LINE SPECTRA

- if you take element and ATOMIZE it, if excited by energy it will emit light at unique frequencies. The set of emitted frequencies is called an ATOMIC LINE SPECTRUM.

- Atomized elements will also ABSORB these same frequencies (but not others)!



... so, why don't atoms by themselves emit continuous spectra like a flashlight would?