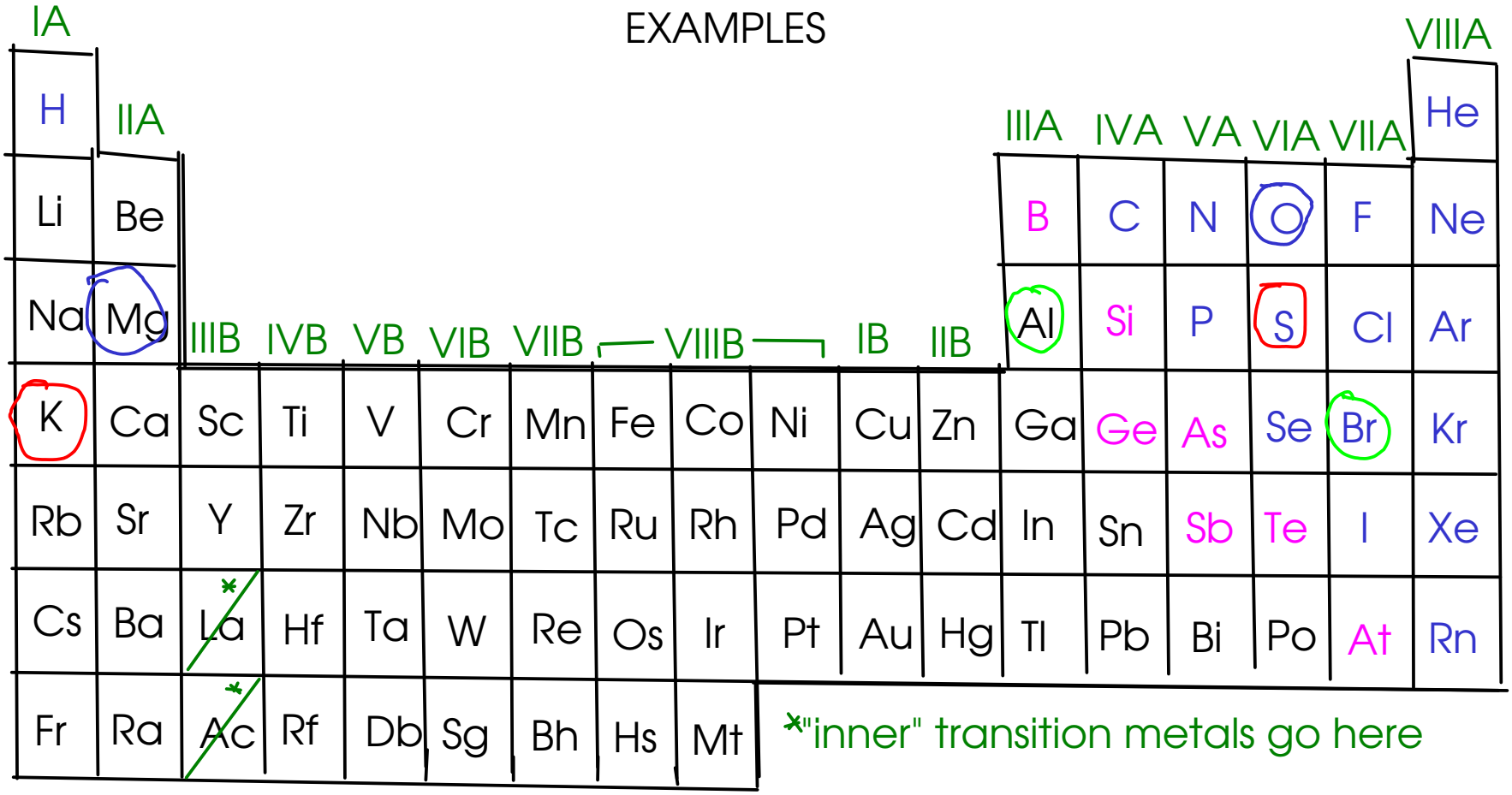
EXAMPLES



Find the formulas of:

(1) an ionic compound containing Al and Br

(2) an ionic compound containing Mg and O

(3) an ionic compound containing S and K

Al^{3+} Br^{-}

Mg^{2+} O^{2-}

S^{2-} K^{+}

Find the formula of:
* an ionic compound containing Al and Br

Al^{3+} Br^-
Br^-
Br^-

$AlBr_3$

Find the formula of:
* an ionic compound containing Mg and O

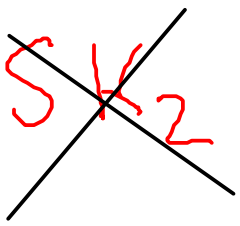
Mg^{2+} O^{2-}

MgO

Find the formula of:
* an ionic compound containing S and K

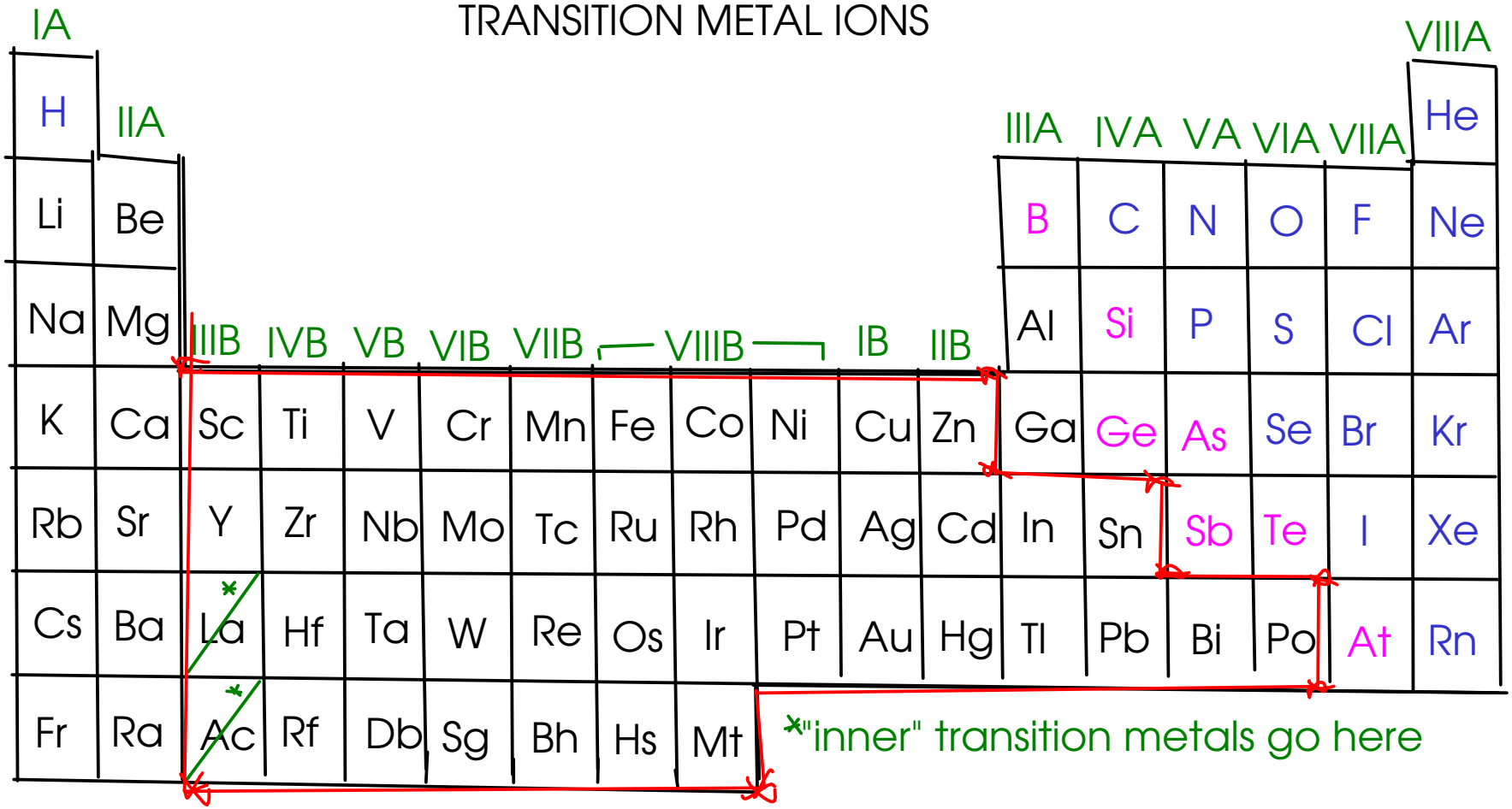
S^{2-} K^+
K^+

K_2S



Remember: Cations (+ charge) go first in ionic formulas.

TRANSITION METAL IONS



The transition metals always form CATIONS!

However, many transition metals are capable of forming SEVERAL DIFFERENT CATIONS!

Example: Iron (Fe) forms two cations, depending on the situation: Fe^{2+} or Fe^{3+}

TRANSITION METAL CATIONS

- So how do you know which cation you're dealing with? For now, you'll have to be told

- Either the chemical formula of an ionic compound or the name of an ionic compound can tell you what charge is on the transition metal cation.

Examples:

Fe_3N_2

Fe^{2+} N^{3-}

Fe^{2+} N^{3-}

Fe^{2+}

$+6$ -6

The iron ions in this compound have a charge of +2 ... and are called "iron(II)" - pronounced "iron two". The compound is called "iron(II) nitride".

FeN

Fe^{3+} N^{3-}

$+3$ -3

The iron ion in this compound has a charge of +3 ... and is called "iron(III)" - pronounced "iron three". The compound is called "iron(III) nitride".

POLYATOMIC IONS

- Some MOLECULES can gain or lose electrons to form CATIONS or ANIONS. These are called POLYATOMIC IONS

- Polyatomic ions form ionic compounds in the same way that single-element ions do.

Example: CO_3^{2-} : carbonate ion

Compare these formulas!

Al^{3+} O^{2-}
Al^{3+} O^{2-}
O^{2-}

Al_2O_3

Al^{3+} Al^{3+} CO_3^{2-} CO_3^{2-} CO_3^{2-}

$Al_2(CO_3)_3$

* Use parenthesis when an ionic compound's formula contains more than one of a polyatomic ion.

A chart of common polyatomic ions is available on the course web site!

(p130 - 7th edition)

NAMES OF IONS

- To properly discuss ions and ionic compounds, we have to know how to name them!

CATIONS

3 kinds:

1. Main group cations (metals that take only one charge when forming ions)

   - The element's name is the same as the ion's name!

   Mg^{2+}: "magnesium ion"

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2. Transition metal cations (from metals that can form several cations)

   - The CHARGE of the cation must be given. Use a ROMAN NUMERAL after the element name to indicate charge!

   Fe^{2+}: "iron(II) ion"

   Cu^{+}: "copper(I) ion"

   Fe^{3+}: "Iron(III) ion"

---

3. Polyatomic cations

   - Memorize list.

   NH_4^+: "ammonium ion"

ANIONS

2 kinds

1. Main-group nonmetals

- Use the STEM NAME of the element, then add "-ide" suffix

N^{3-} : "nitride" ion P^{3-} : "phosphide ion" S^{2-} : sulfide ion

O^{2-} : "oxide ion" F^{-} : "fluoride ion"

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2. Polyatomic ions

- List (see web site) (also p130 in textbook 7th ed)

$C_2H_3O_2^{-}$: "acetate ion" SO_4^{2-} : "sulfate ion"

NO_3^{-} : "nitrate ion" SO_3^{2-} "sulfite ion"

NO_2^{-} : "nitrite ion"

* Polyatomic ions ending in "-ate" and "-ite" suffixes always contain oxygen! "-ate" ions have more oxygen atoms than their "-ite" counterparts.

NAMING IONIC COMPOUNDS

- The name of the compound is based on the name of the ions in the compound
- Cation first, anion second (drop the word "ion")

Examples:

$Mg(OH)_2$

magnesium hydroxide

Na_2S

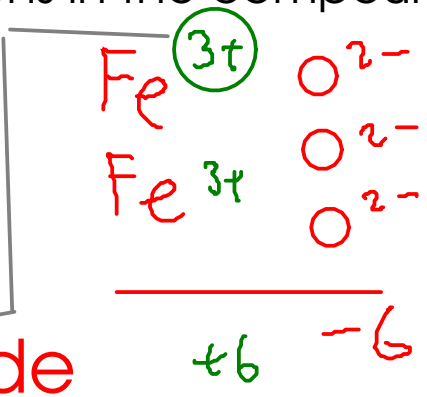
sodium sufide

$BeBr_2$

beryllium bromide

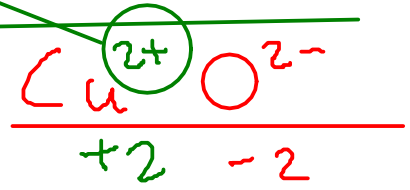
Fe_2O_3

iron(III) oxide



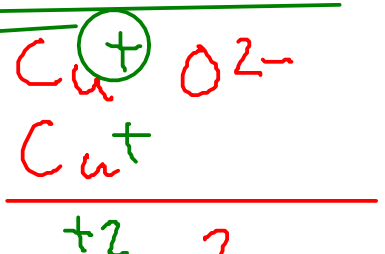
CuO

copper(II) oxide



Cu_2O

copper(I) oxide



Remember to include the Roman numeral for CHARGE in the name of transition metal compounds!