# ACIDS

## (2) OXYACIDS

- Easy to think about as HYDROGEN IONS combined with POLYATOMIC IONS

- These acids are not true ionic compounds, but they interact with water to PRODUCE ions!

- named based on the polyatomic ion they contain, with an ending change:

(1) - ions ending in -ATE form acids ending in -IC

(2) - ions ending in -ITE form acids ending in -OUS

$H_2SO_4$ sulfuric acid

$H_3PO_4$ phosphoric acid

$H_2SO_3$ sulfurous acid

$HNO_3$ nitric acid

# OXYACID EXAMPLES

acetic acid

based on ACETATE ion

$H^+ \quad C_2H_3O_2^-$

$HC_2H_3O_2$

nitrous acid

based on NITRITE ion

$H^+ \quad NO_2^-$

$HNO_2$

carbonic acid

based on CARBONATE ion

$H^+ \quad CO_3^{2-}$

$H^+$

$H_2CO_3$

* The number of hydrogen ions to add to the polyatomic to make the acid equals the charge of the polyatomic.

# SUMMING UP CHEMICAL NOMENCLATURE

- You need to be able to tell, by looking at a name OR a formula, what kind of compound you are working with!

DON'T GET THE NAMING SYSTEMS MIXED UP! EACH KIND OF COMPOUND IS NAMED WITH ITS OWN SYSTEM!

## FROM A CHEMICAL NAME

- If the name has a Roman numeral, the name of a metal, or "ammonium", the compound is likely IONIC

- If the name has a Greek prefix, the compound is BINARY MOLECULAR

- If the name contains the word "acid":

... and starts with "hydro-", then the compound is a BINARY ACID

... and does not start with "hydro-", the compound is an OXYACID

# FROM A CHEMICAL FORMULA

- if the formula contains a metal or the $NH_4^+$ ion, it is likely IONIC

$H_2O$ $H_2O_2$

- If the formula starts with H and is not either water or hydrogen peroxide, the compound is likely an ACID. Which kind?

  - BINARY ACIDS contain only two elements

  - OXYACIDS contains oxygen

- If the formula contains only nonmetals (and is not an ammonium compound or an acid), the compound is likely MOLECULAR

Examples:

$PCl_3$: BINARY MOLECULAR
Name: phosphorus trichloride

$NH_4Cl$: IONIC (ammonium ion)
Name: ammonium chloride

$H_3PO_4$: OXYACID (hydrogen, phosphate)
Name: phosphoric acid

$Fe_2(SO_4)_3$: IONIC (iron - metal!)
Name: iron(III) sulfate

$Fe^{3+}$ $SO_4^{2-}$
$Fe^{3+}$ $SO_4^{2-}$
$SO_4^{2-}$

END OF MATERIAL FOR TEST #2

# CHEMICAL EQUATIONS

- are the "recipes" in chemistry
- show the substances going into a reaction, substances coming out of the reaction, and give other information about the process

"yields"

$$MgCl_2(aq) + 2AgNO_3(aq) \longrightarrow 2AgCl(s) + Mg(NO_3)_2(aq)$$

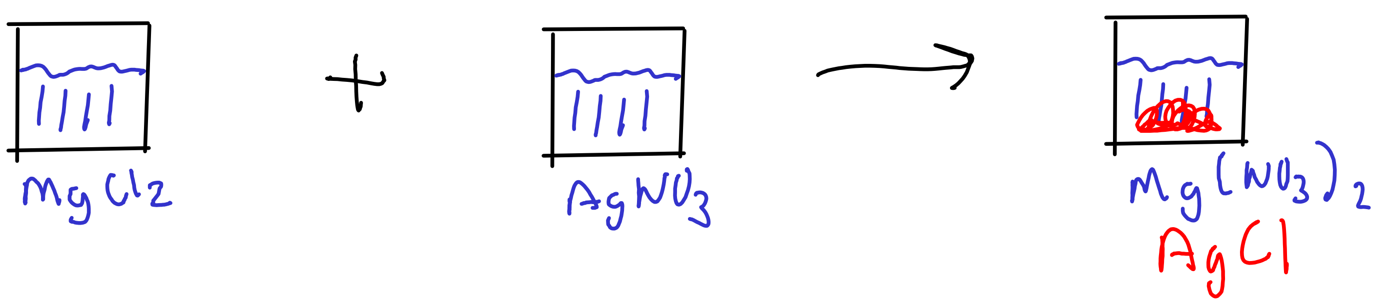
REACTANTS - materials that are needed for a reaction

PRODUCTS - materials that are formed in a reaction

COEFFICIENTS - give the ratio of molecules/atoms of one substance to the others

PHASE LABELS - give the physical state of a substance:

(s) -solid
(l) - liquid
(g) - gas
(aq) - aqueous. In other words, dissolved in water



# CHEMICAL EQUATIONS

$$2\,Mg(s) + O_2(g) \xrightarrow{\Delta} 2\,MgO(s)$$

REACTION CONDITIONS - give conditions necessary for chemical reaction to occur. May be:

- $\Delta$ apply heat
- catalysts - substances that will help reaction proceed faster
- other conditions, such as required temperatures

- Reaction conditions are usually written above the arrow, but may also be written below if the reaction requires several steps or several different conditions

# COEFFICIENTS

- Experimentally, we can usually determine the reactants and products of a reaction

- We can determine the proper ratios of reactants and products WITHOUT further experiments, using a process called BALANCING

- BALANCING a chemical equation is making sure the same number of atoms of each element go into a reaction as come out of it.

- A properly balanced chemical equation has the smallest whole number ratio of reactants and products.

- There are several ways to do this, but we will use a modified trial-and-error procedure.

# BALANCING

$$C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$$ ✓

6      4

10

1. Pick an element. Avoid (if possible) elements that appear in more than one substance on each side of the equation.

2. Change the coefficients on substances containing this element so that the same number of atoms of the element are present on each side. CHANGE AS LITTLE AS POSSIBLE!

3. Repeat 1-2 until all elements are done.

4. Go back and quickly VERIFY that you have the same number of atoms of each element on each side, If you used any fractional coefficients, multiply each coefficient by the DENOMINATOR of your fraction.

Use SMALLEST WHOLE NUMBER RATIOS!

# BALANCING

$$3MgCl_2 + 2Na_3PO_4 \longrightarrow Mg_3(PO_4)_2 + 6NaCl$$

$$C_2H_2 + \frac{5}{2}O_2 \longrightarrow 2CO_2 + H_2O$$

$$\underline{4 \qquad\qquad 1}$$
$$5$$

... to get a SINGLE oxygen atom from molecular oxygen, we need HALF of an oxygen molecule. To get FIVE oxygen atoms, we need FIVE HALVES of oxygen molecule.

To get rid of the fraction, multiple ALL coefficients by the denominator of the fraction (in this case it's 2):

$$2C_2H_2 + 5O_2 \longrightarrow 4CO_2 + 2H_2O$$

$$H_2SO_4 + 2NaOH \longrightarrow Na_2SO_4 + 2H_2O$$

H: 2 + 2 = 4      H: ~~2~~ 4

O: 4 + 2 = 6      O: 4 + 2 = 6

Initially, we skip H and O because they appear in more than one substance on each side of the equation.