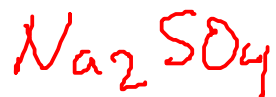
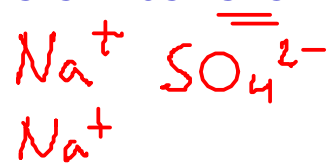
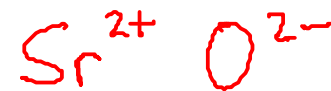


DETERMINING IONIC FORMULAS

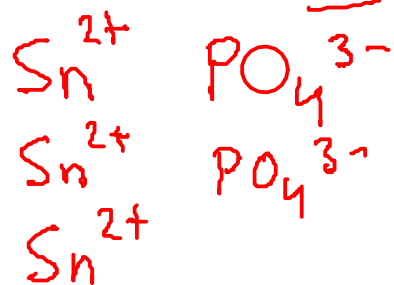
sodium sulfate



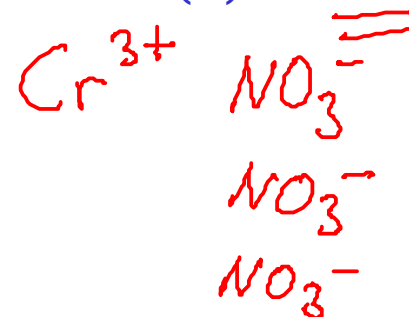
strontium oxide



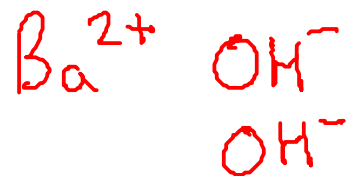
tin(II) phosphate



chromium(III) nitrate

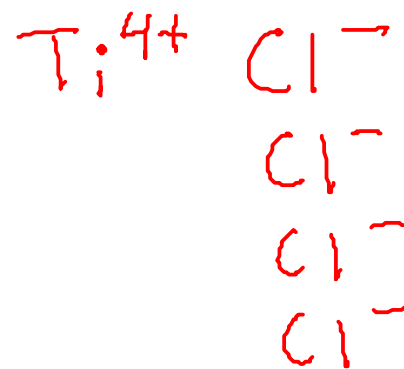


barium hydroxide



Don't forget the parenthesis when you have more than one hydroxide ion!

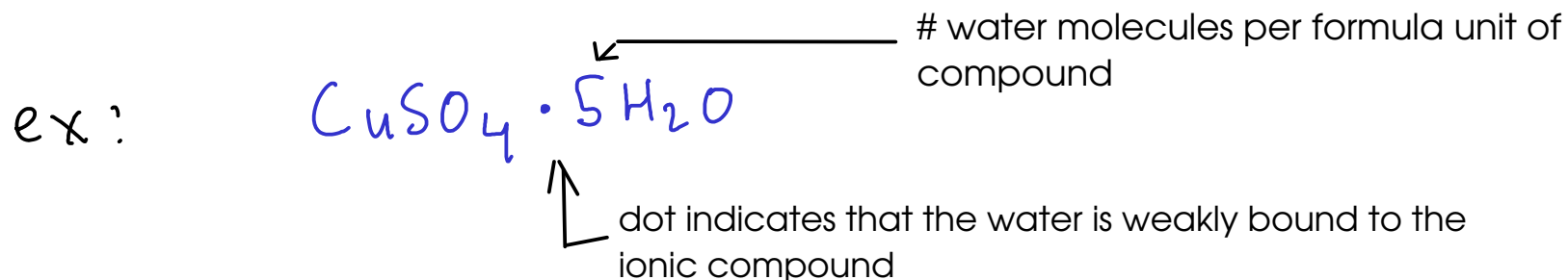
titanium(IV) chloride



HYDRATES

- many ionic compounds are formed by crystallizing the compound from water. Sometimes, this causes water molecules to become part of the crystal structure.

- This water is present in a definite ratio to the ions in the compound. Can be removed by heating, but will NOT evaporate if the compound is left standing.



- many DESSICANTS are hydrates that have had their water molecules driven off. They will slowly reabsorb water from the air (and keep the environment in a dessicator at a low humidity)

- Hydrates are named using the name of the ionic compound, and a Greek prefix in front of the word "hydrate" to indicate how many water molecules are associated

