

NAMING IONIC COMPOUNDS

- The name of the compound is based on the name of the ions in the compound
- Cation first, anion second (drop the word "ion")

Examples:



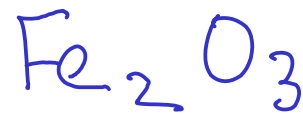
magnesium hydroxide



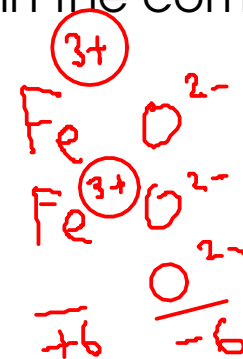
sodium sulfide



beryllium bromide



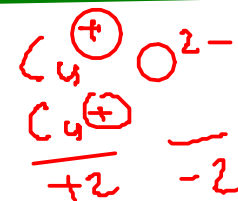
iron(III) oxide



copper(II) oxide

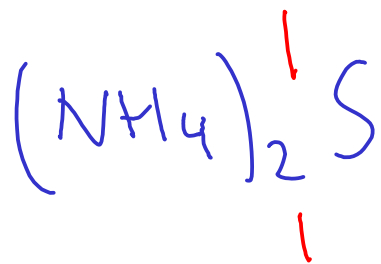


copper(I) oxide

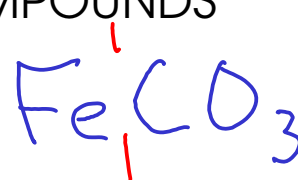


Remember to include the Roman numeral for CHARGE in the name of transition metal compounds!

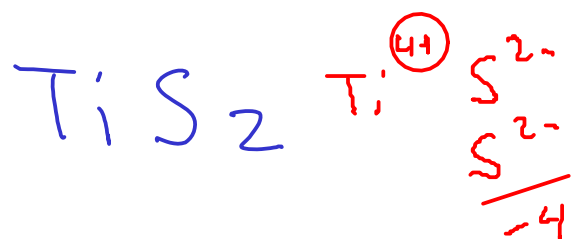
NAMING IONIC COMPOUNDS



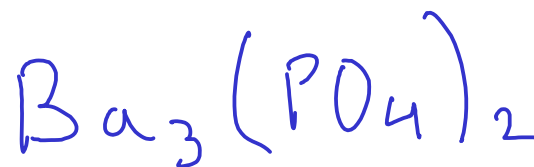
ammonium sulfide



iron(II) carbonate



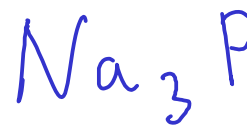
titanium(IV) sulfide



barium phosphate



calcium nitrate



sodium phosphide

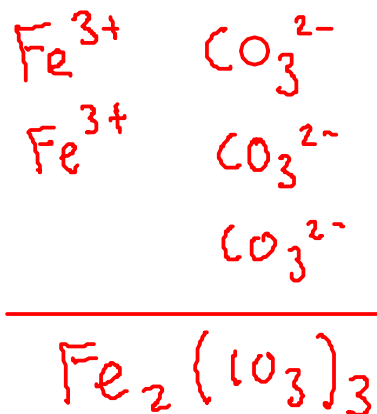
Spelling matters!

DETERMINING THE FORMULA OF AN IONIC COMPOUND FROM THE NAME

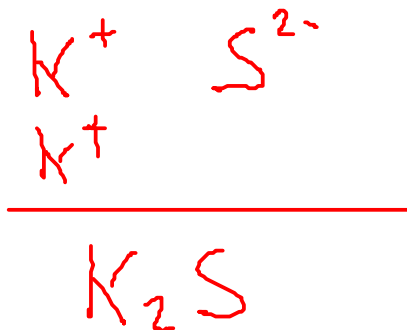
- The name of an ionic compound is made of the names of the CATION and ANION in the compound.
 - To get the FORMULA, you must figure out the SMALLEST RATIO of cation to anion that makes the charges balance out
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Examples:

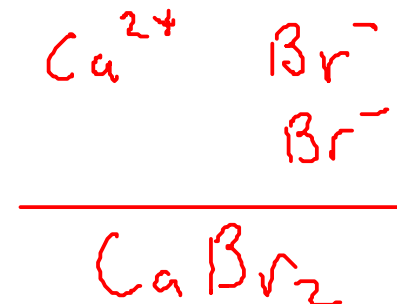
iron(III) carbonate



potassium sulfide

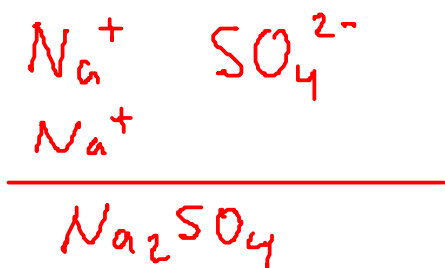


calcium bromide

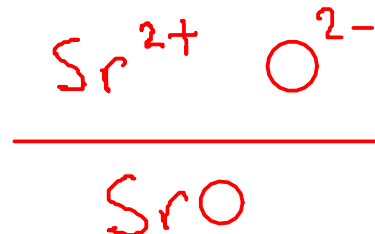


DETERMINING IONIC FORMULAS

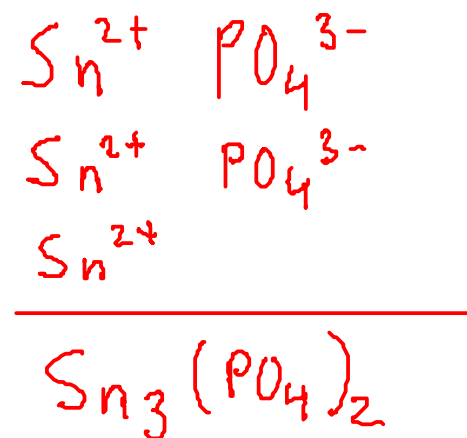
sodium sulfate



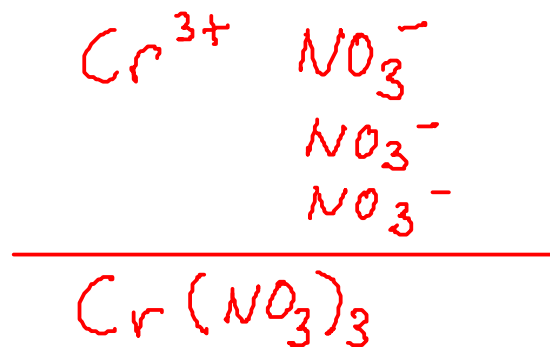
strontium oxide



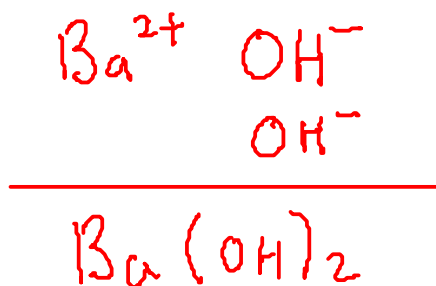
tin(II) phosphate



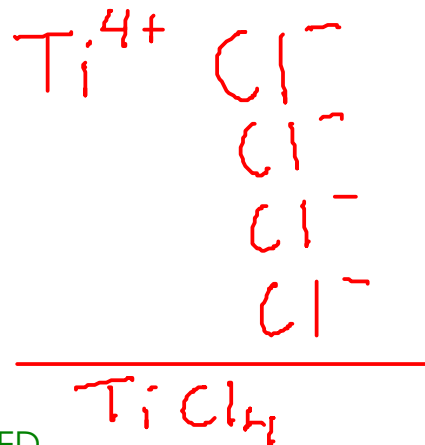
chromium(III) nitrate



barium hydroxide



titanium(IV) chloride



~~BaOH_2~~ If your formula contains more than one polyatomic ion, you NEED to use parenthesis!