Measurements

Measurements are comparisons of properties against accepted standards, called units.

English/US units:

1 foot = 12 inches 1 yard = 3 feet 1 mile = 1760 yards

5280 feet = 1 mile

So what's the problem?

English units are not consistent. This makes the English system difficult to learn and use reliably. The relationships between every unit in English must me memorized separately.

English units are nonstandard and difficult to use. Solution?

THE METRIC SYSTEM

Metric Base Units:

Length	meter	m
Mass	kilogram *	kg
Temperature	Kelvin	K
Time	second	s

All metric units are made up of COMBINATIONS of BASE UNITS!

* we usually treat the gram as if it's the base unit for mass!

- One meter is approximately 3.3 feet.
- One kilogram is approximately 2.2 pounds.

What about SIZE?

Metric units may be made larger or smaller by adding PREFIXES.

Metric Prefixes:

mega-	10^{6}	M
kilo-	10^{3}	k
centi-	10^{-2}	c
milli-	10^{-3}	m
micro-	10^{-6}	μ

Bigger units

smaller units

Memorize these prefixes!

Applying prefixes

1 _____ m = _____ m

1 km = 10^{3} m (1000 m) $10 \times 10 \times 10$

1 cm = 10^{-2} m ($\frac{1}{100}$ m) $\frac{1}{10} \times \frac{1}{10}$

Scaling units with metric prefixes ... examples

The distance between here and Columbia, SC is about 107,000 meters. What metric unit would be best suited for a distance like this?

$km = 10^3 m = 1000 m$

107 km

A piece of chalk is 0.080 meters long. What metric unit would be best suited for this length?

$c = 10^{-2}$ $\quad cm = 10^{-2} m = \frac{1}{100} m$

8.0 cm

Derived Units

- are units that are made up of combinations of metric base units with each other and/or with prefixes

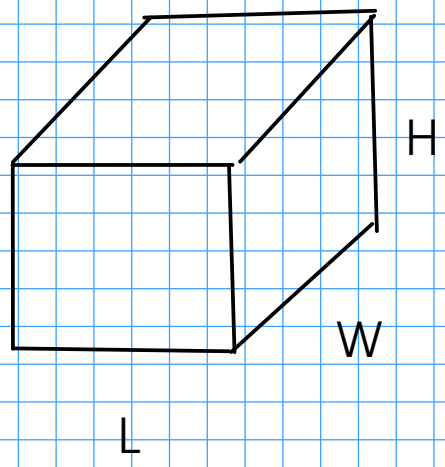
Example: speed $\frac{\text{miles}}{\text{hr}}$, $\frac{\text{Km}}{\text{hr}}$ $\left(\frac{\text{length}}{\text{time}}\right)$ $\frac{\text{m}}{\text{s}}$

Two derived units are particularly important in introductory chemistry:

1) VOLUME

2) DENSITY

VOLUME



VOLUME = L × W × H

What are the units of volume in the metric system?

L = length ; metric base unit for length is: meter

W = meters H = meters

$$\text{VOLUME} = \underset{L}{(\text{m})} \times \underset{W}{(\text{m})} \times \underset{H}{(\text{m})} = \text{m}^3 \quad \text{"cubic meters"}$$

... but the cubic meter is a very large unit (when you're talking about a laboratory or a medical setting). It's a cube that's a meter (about a yard) long on every side. We would prefer a smaller unit.

Practical issues for volume units

- Cubic meters are too large! A meter is very similar in length to a yard, so a cubic meter is a cube that is approximately a yard long on each side!

A smaller unit for volume?

Cubic decimeters! dm^3

(decimeter = 1/10 meter)

Cubic decimeters are given the name "liters", abbreviation "L"

In the lab, we typically need an even smaller unit than the liter, so we use milliliters (mL)

"cc"
cubic centimeter
=
milliliter

$1\ mL = 10^{-3}\ L$

-or-

$1000\ mL = 1\ L$

DENSITY

- Density is a measure of the concentration of matter; of how much matter is present in a given space

- Density is defined as the MASS per unit VOLUME, or ...

$$\text{Density} = \frac{\text{mass}}{\text{Volume}}$$

What are the metric units of DENSITY?

mass : kg

volume : m^3

So, density unit = $\boxed{\frac{kg}{m^3}}$

... these units (kg and cubic meters) are large for lab work.

In the lab, we typically measure masses as grams and volumes as milliliters, so the density unit we will use most often is:

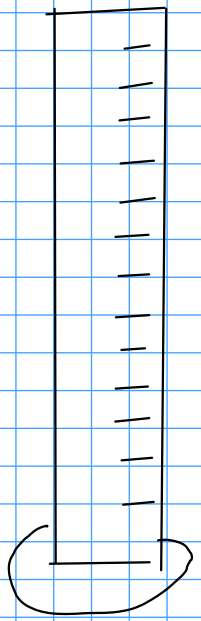
$$\frac{g}{mL}$$

A useful density to remember:

WATER at room temp: Density = 1 g/mL

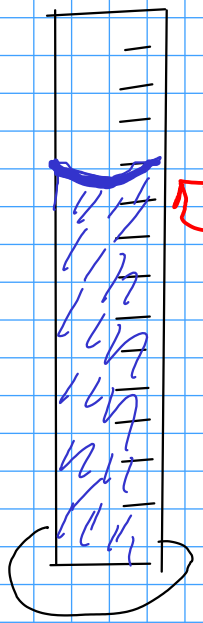
Measuring density

... of a liquid



1) Measure mass of empty cylinder

mass = 97.35 g



2) Fill cylinder and measure volume of liquid

Volume = 25.3 mL

3) Measure mass of filled cylinder

mass = 130.55 g

4) Subtract to find mass of liquid

$$\begin{array}{rr} 130.55 & g \\ - \ 97.35 & g \\ \hline 33.20 & g \end{array}$$

5) Density = mass liquid / volume liquid

$$\text{Density} = \frac{33.20 \ g}{25.3 \ mL}$$

$$= \boxed{1.31 \ g/mL}$$