BINARY MOLECULAR COMPOUNDS

Examples:

BF_3 — boron trifluoride

Cl_2O_7 — dichlorine hept(a)oxide

CO — carbon monoxide

CO_2 — carbon dioxide

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carbon tetrachloride — CCl_4

dihydrogen monoxide — H_2O

dinitrogen tetrafluoride — N_2F_4

$MgCl_2$: "magnesium chloride", NOT "magnesium dichloride". Why not? This compound contains a METAL and a nonmetal. It's IONIC.

ACIDS

1 BINARY ACIDS

- named after the element (other than hydrogen) they contain
- common binary acids include a Group VIIA element
- named: "Hydro-" + STEM NAME OF ELEMENT+ "-ic acid"

Four common binary acids

- HF : hydrofluoric acid * dissolves glass!
- HCl : hydrochloric acid * most common binary acid!
- HBr : hydrobromic acid
- HI : hydroiodic acid

ACIDS

(2) OXYACIDS

- Easy to think about as HYDROGEN IONS combined with POLYATOMIC IONS

- These acids are not true ionic compounds, but they interact with water to PRODUCE ions!

- named based on the polyatomic ion they contain, with an ending change:

(1) - ions ending in -ATE form acids ending in -IC

(2) - ions ending in -ITE form acids ending in -OUS

H_2SO_4 sulfuric acid

H_3PO_4 phosphoric acid

H_2SO_3 sulfurous acid

HNO_3 nitric acid

H_2CO_3 carbonic acid

OXYACID EXAMPLES

acetic acid

$$H^+ \quad C_2H_3O_2^-$$

$$HC_2H_3O_2$$

nitrous acid

$$H^+ \quad NO_2^-$$

$$HNO_2$$

carbonic acid

$$H^+ \quad CO_3^{2-}$$

$$H^+$$

$$H_2CO_3$$

SUMMING UP CHEMICAL NOMENCLATURE

- You need to be able to tell, by looking at a name OR a formula, what kind of compound you are working with!

DON'T GET THE NAMING SYSTEMS MIXED UP! EACH KIND OF COMPOUND IS NAMED WITH ITS OWN SYSTEM!

FROM A CHEMICAL NAME

- If the name has a Roman numeral, the name of a metal, or "ammonium", the compound is likely IONIC

- If the name has a Greek prefix, the compound is BINARY MOLECULAR

- If the name contains the word "acid":

  ... and starts with "hydro-", then the compound is a BINARY ACID

  ... and does not start with "hydro-", the compound is an OXYACID

FROM A CHEMICAL FORMULA

- if the formula contains a metal or the NH_4^+ ion, it is likely IONIC

- If the formula starts with H and is not either water or hydrogen peroxide, the compiund is likely an ACID. Which kind?

  - BINARY ACIDS contain only two elements

  - OXYACIDS contains oxygen

- If the formula contains only nonmetals (and is not an ammonium compound or an acid), the compound is likely MOLECULAR

Examples:

PCl_3 : BINARY MOLECULAR
Name: phosphorus trichloride

NH_4Cl : IONIC (ammonium ion)
Name: ammonium chloride

H_3PO_4 : OXYACID (hydrogen, phosphate)
Name: phosphoric acid

$Fe_2(SO_4)_3$: IONIC (iron - metal!)
Name: iron(III) sulfate

Fe^{3+} SO_4^{2-}
Fe^{3+} SO_4^{2-}
SO_4^{2-}