

Blocks on the periodic table

11
Na
Sodium
22.99

Atomic number: This is always a whole number. The periodic table is arranged by atomic number!

Element symbol: A one or two letter abbreviation for the name of the element. Sometimes, the abbreviation is based on a language OTHER THAN ENGLISH! (Example: Na is short for "natrium", the Latin name of sodium.)

Element name: Sometimes, this is left off of periodic tables, especially small ones!

Atomic weight: This is either a decimal number or a number in parenthesis.

88
Ra
Radium
(226)

For RADIOACTIVE ELEMENTS - elements where the atomic nucleus breaks down, causing the atom to break apart - the MASS NUMBER of the most stable ISOTOPE is given in (parenthesis) instead of the atomic number!

CHEMICAL COMPOUNDS

- Dalton's theory does not mention this, but there is more than one way for atoms to come together to make chemical compounds!
- There are TWO common kinds of chemical compound, classified based on how the atoms in the compound are held together:

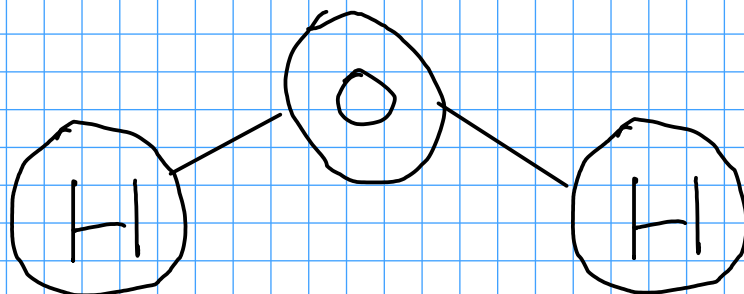
① MOLECULAR COMPOUNDS

② IONIC COMPOUNDS

MOLECULAR COMPOUNDS

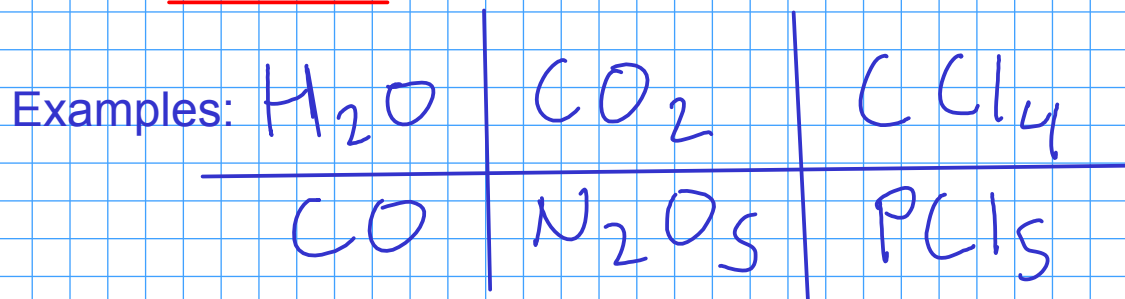
"covalent bonds"

- form when atoms SHARE outer electrons with each other. This results in a set of connected atoms called a MOLECULE



Stick figure of a water (H_2O) molecule

- usually form between nonmetals and other nonmetals or between nonmetals and metalloids



CANDLE WAX is made up of molecular compounds

- some solid at room temperature. These solids tend to have low melting points.

PCl_5 is a solid, $mp = 180^\circ C$

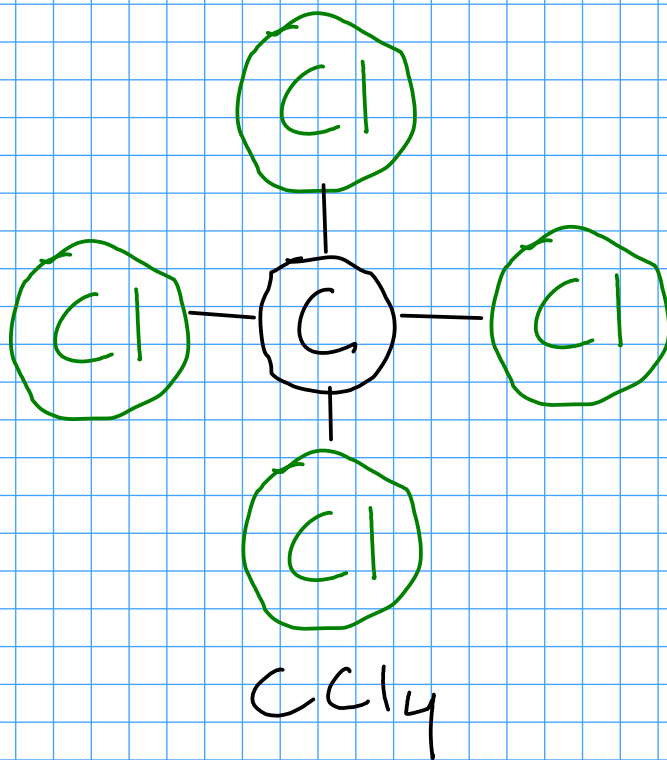
- many are liquids or gases at room temperature

H_2O, CCl_4 = liquids CO, CO_2, N_2O_5 = gases

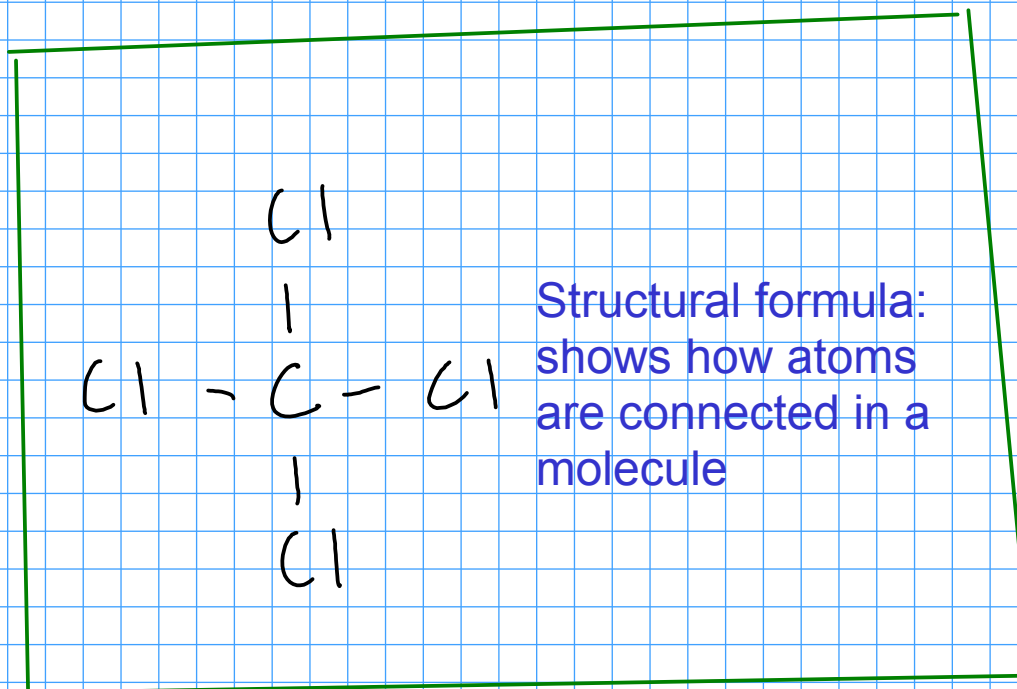
MOLECULAR FORMULAS

- formula of a molecular compound represents the EXACT NUMBER OF ATOMS OF EACH ELEMENT in a single molecule of the compound

Example: Each molecule of CCl_4 contains exactly one carbon atom and four chlorine atoms



"ball and stick" model



IONIC COMPOUNDS

- formed when atoms TRANSFER ELECTRONS between each other forming charged atoms, called IONS.

Two kinds of ions:

cation

① CATIONS: formed when an atom **LOSES** one or more electrons.

- overall, a cation has a **POSITIVE** charge, because it has more protons in the nucleus than electrons in the electron cloud

- usually formed by METALS, but occasionally hydrogen will also form a cation

② ANIONS: formed when an atom **GAINS** one or more electrons

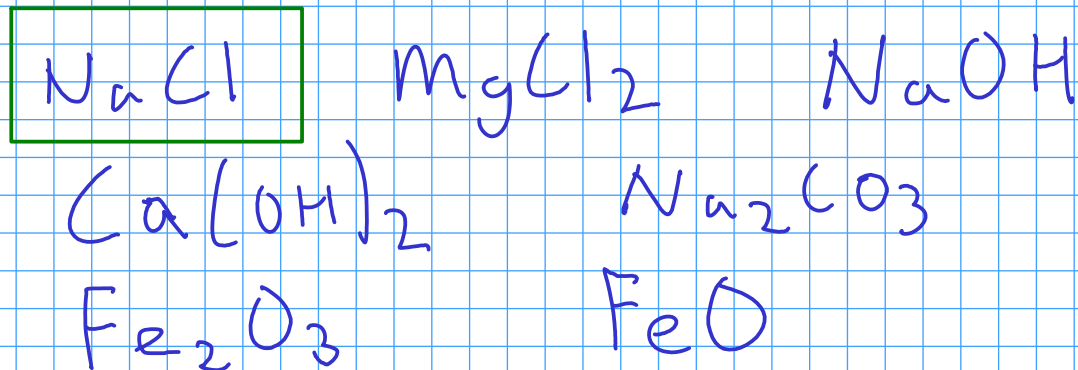
- overall, an anion has a **NEGATIVE** charge, because it has more electrons in the electron cloud than protons in the nucleus

- usually formed by NONMETALS

IONIC COMPOUNDS

- USUALLY form from metals combining with nonmetals, or from metals combining with metalloids

Examples:



- almost always solid at room temperature, and usually have relatively high melting points

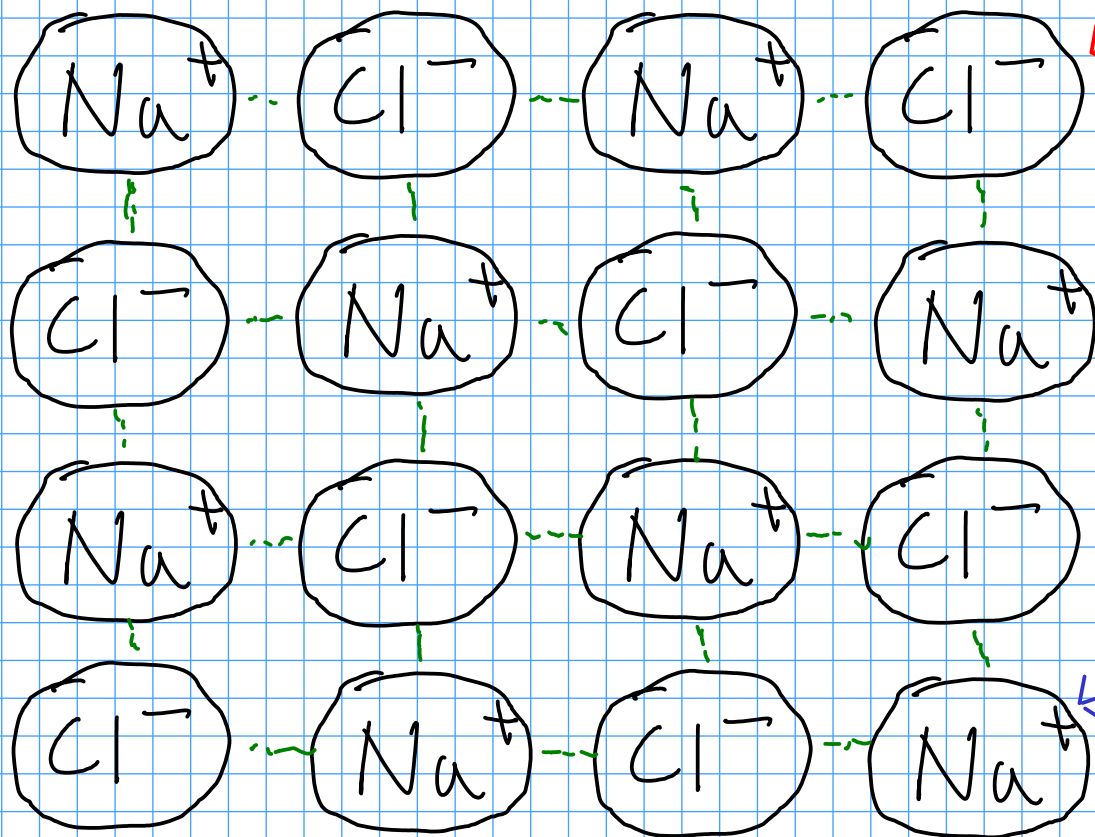
All of the above are solids at room temperature. NaCl has a melting point of 801°C .

- as solids, do not conduct electricity. If dissolved in water (some do not dissolve significantly in water), will form a solution that conducts electricity.

IONIC COMPOUNDS

- ionic compounds are held together by ELECTROSTATIC INTERACTIONS

(in other words, the attraction between oppositely charged ions!)



Each chloride ion is strongly attracted to ALL of the sodium ions surrounding it!

Each sodium ion is strongly attracted to ALL of the chlorine atoms surrounding it!

There are no "molecules" in ionic compounds - in the sense that you can't point to a discrete unit of atoms that are connected to only each other

IONIC FORMULAS

- since there are no "molecules", an ionic formula cannot describe how many and what kinds of atoms are in a molecule!
- all ionic compounds are observed to be (overall) electrically neutral, so the IONS they contain must be present in such a way that the charges **BALANCE EACH OTHER**
- an ionic formula gives the SMALLEST WHOLE NUMBER RATIO OF CATION TO ANION in the ionic compound

