CHM 111
Quick Quiz - 3/15/04 **Name: _____**

Answer the question. [20]

1) The equilibrium *2NOBr(g) <--> 2NO(g) + Br_2(g)* has an equilibrium constant (K_c) value of 3.07×10^{-4} at 24°C. Does this reaction favor products or reactants at equilibrium?

- _____

2) Define the term *chemical equilibrium*.

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Write concentration-based equilibrium constant expressions for the following reactions. [20]

3) $NH_3(aq) + H_2O(l) <--> NH_4^+(aq) + OH^-(aq)$ $K_c =$
4) $Ag_2CO_3(s) <--> 2Ag^+(aq) + CO_3^{2-}(aq)$ $K_c =$
5) $2SO_2(g) + O_2(g) <--> 2SO_3(g)$ $K_c =$