

CHM 111  
Quick Quiz - 3/15/04

Name: \_\_\_\_\_

Answer the question. [20]

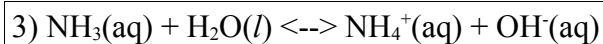
1) The equilibrium  $2NOBr(g) \rightleftharpoons 2NO(g) + Br_2(g)$  has an equilibrium constant ( $K_c$ ) value of  $3.07 \times 10^{-4}$  at  $24^\circ\text{C}$ . Does this reaction favor products or reactants at equilibrium?

• \_\_\_\_\_

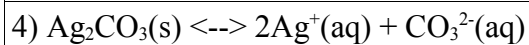
2) Define the term *chemical equilibrium*.

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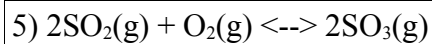
Write concentration-based equilibrium constant expressions for the following reactions. [20]



$K_c =$



$K_c =$



$K_c =$