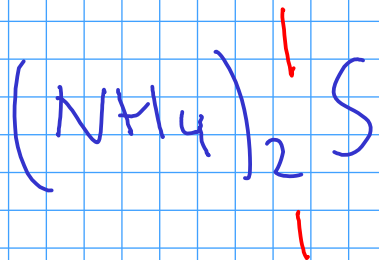
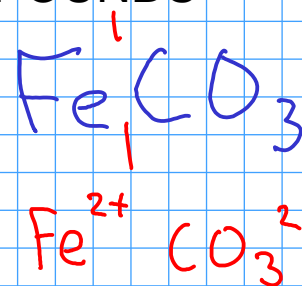


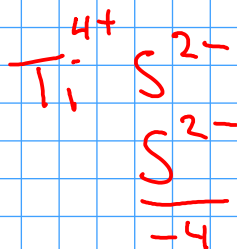
NAMING IONIC COMPOUNDS



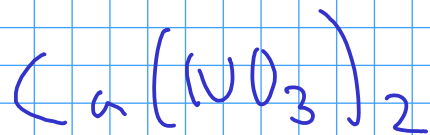
ammonium sulfide



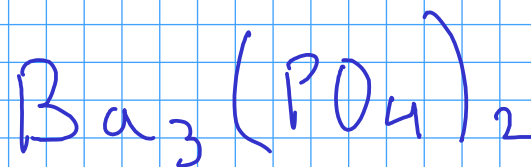
iron(II) carbonate



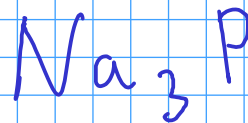
titanium(IV) sulfide



calcium nitrate



barium phosphate



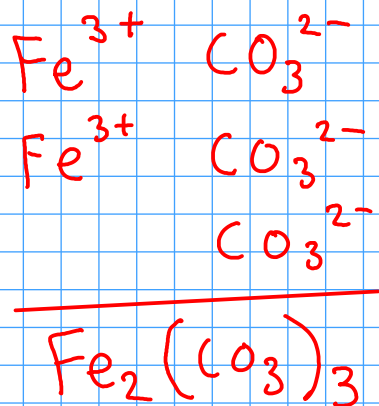
sodium phosphide

DETERMINING THE FORMULA OF AN IONIC COMPOUND FROM THE NAME

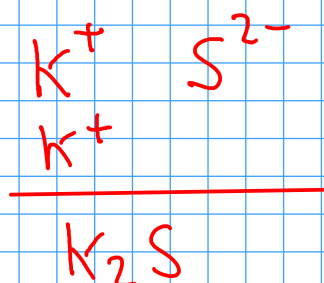
- The name of an ionic compound is made of the names of the CATION and ANION in the compound.
 - To get the FORMULA, you must figure out the SMALLEST RATIO of cation to anion that makes the charges balance out
-

Examples:

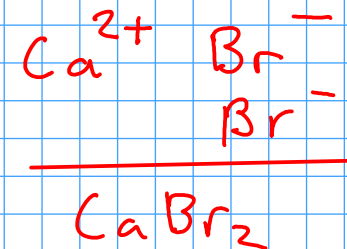
iron(III) carbonate



potassium sulfide

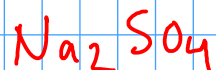
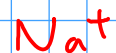
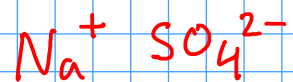


calcium bromide



DETERMINING IONIC FORMULAS

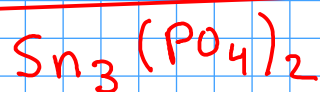
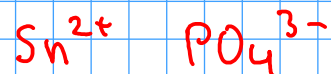
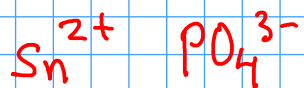
sodium sulfate



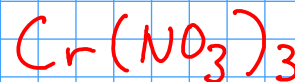
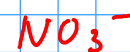
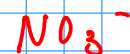
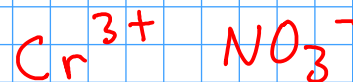
strontium oxide



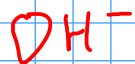
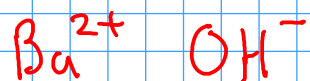
tin(II) phosphate



chromium(III) nitrate



barium hydroxide



titanium(IV) chloride

