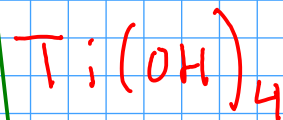
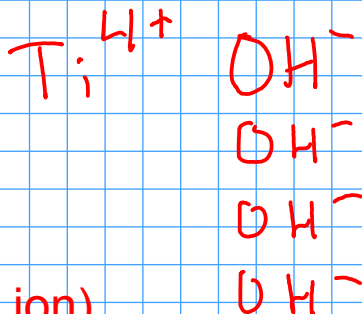


# A REVIEW OF THE CHEMICAL FORMULAS QUIZ

titanium(IV) hydroxide

Ionic!

- presence of metal
- Roman numeral
- "hydroxide" (polyatomic ion)



sulfur hexafluoride

Molecular!

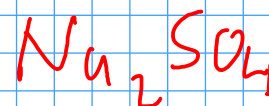
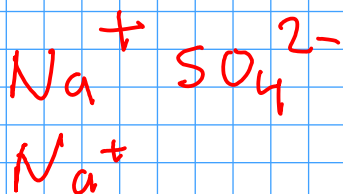
- only nonmetals
- "hexa" prefix



sodium sulfate

Ionic!

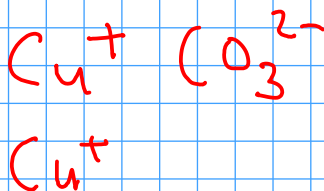
- presence of metal
- "sulfate" (polyatomic ion)



copper(I) carbonate

Ionic!

- presence of metal
- Roman numeral
- "carbonate" (polyatomic ion)

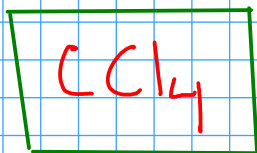


# A REVIEW OF THE CHEMICAL FORMULAS QUIZ

carbon tetrachloride

**Molecular!**

- nonmetals only
- "tetra" prefix



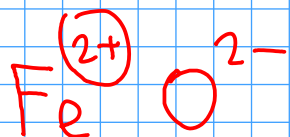
$\text{PCl}_3$

**Molecular!**

- nonmetals only, no polyatomic ions

phosphorus trichloride

$\text{FeO}$



iron(II) oxide

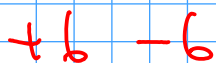
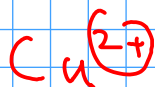
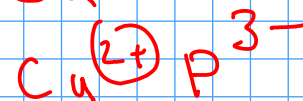
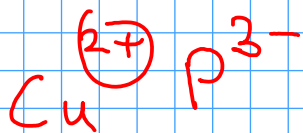
**Ionic!**

- metal (iron)
- note: iron is a transition metal! Find the charge!

$\text{Cu}_3\text{P}_2$

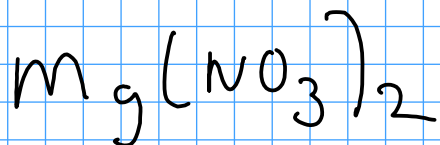
**Ionic!**

- metal (copper)
- note: copper is a transition metal! Find the charge!



copper(II) phosphide

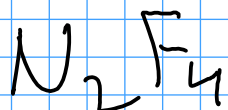
## A REVIEW OF THE CHEMICAL FORMULAS QUIZ



magnesium nitrate

Ionic!

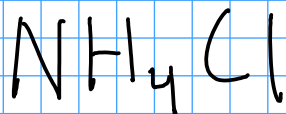
- metal (magnesium)
- polyatomic ion:  $\text{NO}_3^-$



dinitrogen tetrafluoride

Molecular!

- only nonmetals, no polyatomic ions



ammonium chloride

Ionic!

- polyatomic ion:  $\text{NH}_4^+$

This one wasn't on the quiz, but we talked about it as an example of an ionic compound with no metals!

## 10/17/07 TEST

- Covers all of chapters 4 and 5, and parts of chapters 3 and 6
- Use the study guides on the WEB SITE to prepare for the test!
- You will be provided with a periodic table for the test. You will need to bring a calculator and something to write with yourself!

**SOME STUDY TIPS: (These are NOT a substitute for the study guides on the web site! Use those to plan most of your studying!)**

### CHAPTER 3

- Study sections 3.6, 3.7, and 3.8
- Review specific heat calculations and energy unit conversions

### CHAPTER 4

- Study all sections of Chapter 4, but ignore calculations on p 97-98
- Find elements and their information QUICKLY in the periodic table!
- Review differences between molecular and ionic compounds

## 10/17/07 TEST

### CHAPTER 5

- Study all sections of Chapter 5, but your class notes and the flowcharts from the WEB SITE will come in handy
- Review how to determine whether you have an ionic, binary molecular, or acidic compound
- Names to formulas, formulas to names! Review your quizzes and the practice sets on the web

### CHAPTER 6

- Study sections 6.1, 6.2, 6.3, 6.4, 6.5, and 6.10 - but concentrate on your class notes and the chemical reactions handout on the WEB SITE
- Practice balancing chemical equations. It's not hard, but you need to try it before the test!