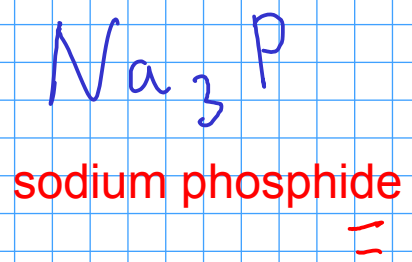
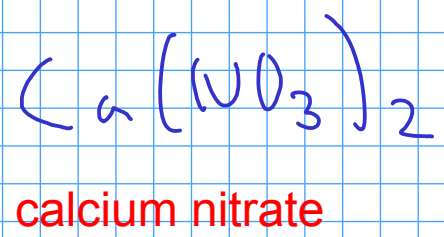
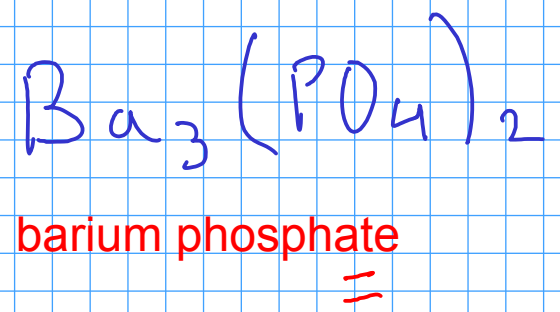
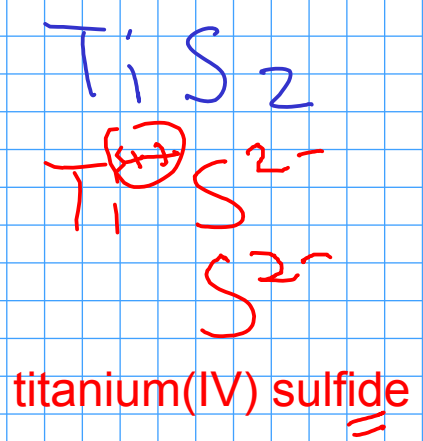
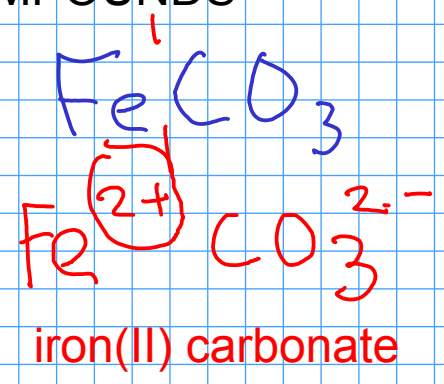
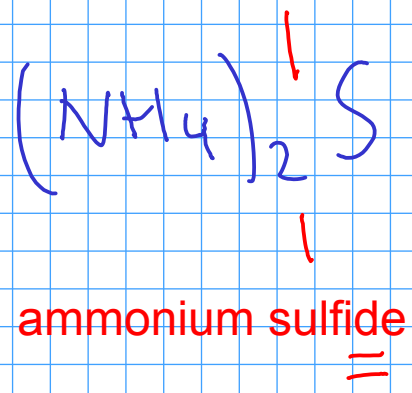


# NAMING IONIC COMPOUNDS

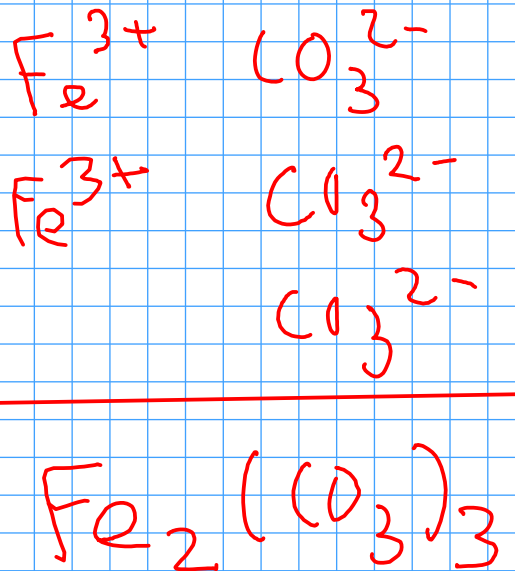


# DETERMINING THE FORMULA OF AN IONIC COMPOUND FROM THE NAME

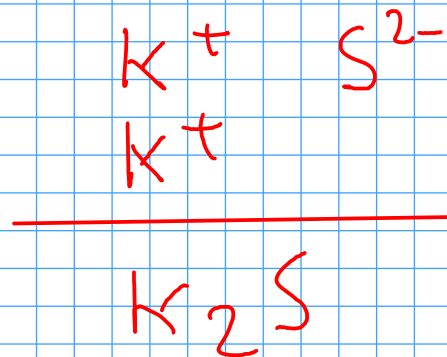
- The name of an ionic compound is made of the names of the CATION and ANION in the compound.
  - To get the FORMULA, you must figure out the SMALLEST RATIO of cation to anion that makes the charges balance out
- 

## Examples:

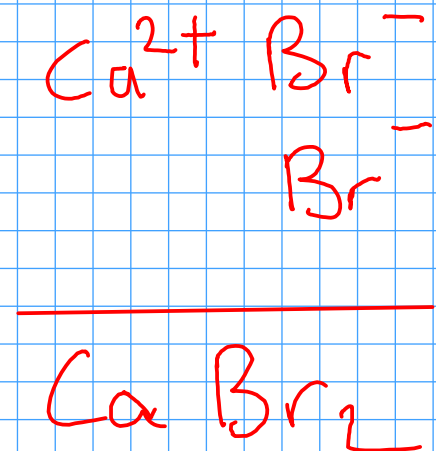
iron(III) carbonate



potassium sulfide

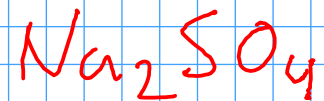
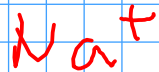
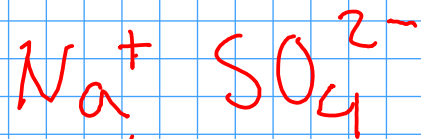


calcium bromide

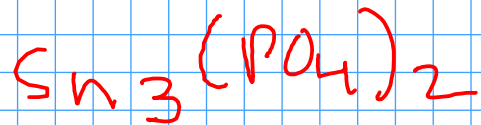
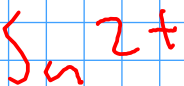
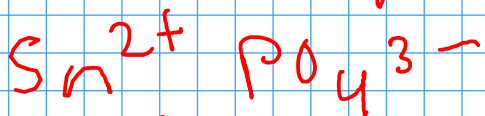
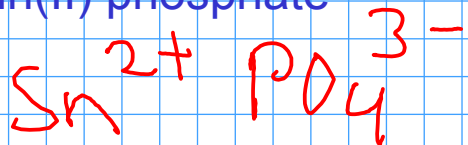


# DETERMINING IONIC FORMULAS

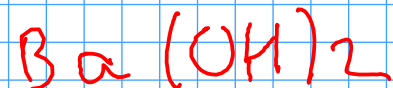
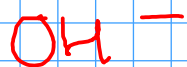
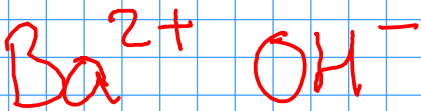
sodium sulfate



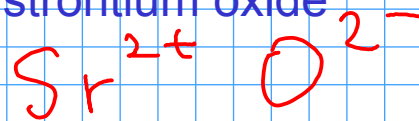
tin(II) phosphate



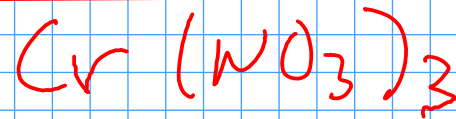
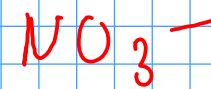
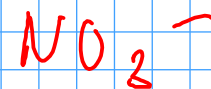
barium hydroxide



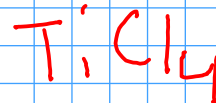
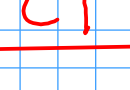
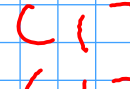
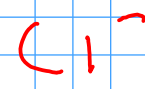
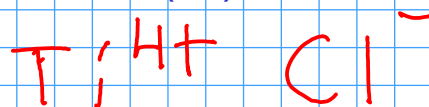
strontium oxide



chromium(III) nitrate



titanium(IV) chloride



# MOLECULAR COMPOUNDS

- There are several kinds of molecular compound. We will learn to name two simple but important classes

## ① BINARY MOLECULAR COMPOUNDS

- molecular compounds containing only two elements

## ② ACIDS

- molecular compounds that dissolve in water to release  $H^+$  ions
- corrosive to metals (react with many to produce hydrogen gas)
- contact hazard: can cause chemical burns to eyes and skin
- sour taste
- turn litmus indicator RED
- two kinds of acids:

### ① BINARY ACIDS

- contain hydrogen and one other element

### ② OXYACIDS

- contain hydrogen, OXYGEN, and another element