

Alchemists

- tried to change "base metals" into gold
- developed many techniques later used in modern chemistry

Modern

Lavosier (1) Made chemistry a quantitative science
(2) nature of combustion
(3) list of elements

Dalton: 1808. "atomic theory"

Measurements

Measurements are comparisons of properties against accepted standards, called units.

English/US units:

$$1 \text{ foot} = 12 \text{ inches} \quad 1 \text{ yard} = 3 \text{ ft}$$

$$1 \text{ mile} = 5280 \text{ ft}$$

$$1760 \text{ yd} = 1 \text{ mile}$$

PROBLEM: Units are not
consistent, hard to use

English units are nonstandard and difficult to use. Solution?

THE METRIC SYSTEM

Base Units:

| | | |
|--------------|-----------|----|
| Length: | meter | m |
| Mass: | kilogram* | kg |
| Time: | second | s |
| Temperature: | Kelvin | K |

All metric units are made from combinations of base units!

*we usually treat the gram as if it's the base unit for mass!

What about size?

Metric units may be made larger or smaller by adding PREFIXES.

| | | |
|--------|-----------|-------|
| mega- | 10^6 | M |
| kilo- | 10^3 | k |
| centi- | 10^{-2} | c |
| milli- | 10^{-3} | m |
| micro- | 10^{-6} | μ |

Bigger units

Smaller units

Memorize
these
prefixes!

Applying prefixes

$$1 \text{ _____ m} = \text{_____ m}$$

$$1 \text{ km} = 10^3 \text{ m} \quad (1000 \text{ m}) \quad 10 \times 10 \times 10$$

$$1 \text{ cm} = 10^{-2} \text{ m} \quad \left(\frac{1}{100} \text{ m} \right) \quad \frac{1}{10} \times \frac{1}{10}$$

Scaling units with metric prefixes ... examples

The distance between here and Columbia, SC is about 107,000 meters.
What metric unit would be best suited for a distance like this?

$$107\,000\text{ m}$$

$$\text{km} = 1000\text{ m},$$

$$\underline{107\text{ km}}$$

A piece of chalk is 0.080 meters long. What metric unit would be best suited for this length?

$$\text{cm} = \frac{1}{100}\text{ m}$$

$$8,0\text{ cm}$$

Derived Units

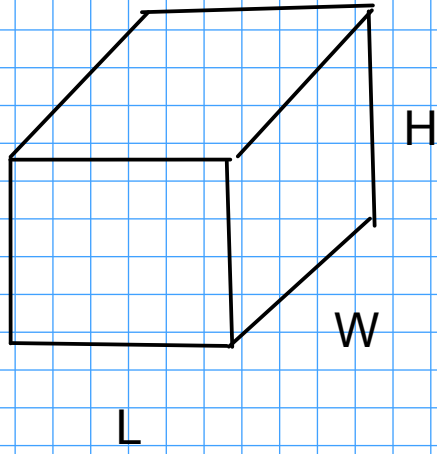
- are units that are made up of combinations of metric base units with each other and/or with prefixes

Two derived units are particularly important in introductory chemistry:

1) VOLUME

2) DENSITY

VOLUME



$$\text{VOLUME} = L \times W \times H$$

UNITS?

metric length unit = meter

So,

$$\begin{aligned} \text{Volume} &= \text{meter} \times \text{meter} \times \text{meter} \\ &= \text{m}^3 \text{ (cubic meter)} \end{aligned}$$

Practical issues for volume units

- Cubic meters are too large! A meter is very similar in length to a yard, so a cubic meter is a cube that is approximately a yard long on each side!

A smaller unit for volume?

Cubic decimeters!

(decimeter = $\frac{1}{10}$ meter)

Cubic decimeters are given the name "liters", abbreviation "L"

In the lab, we typically need an even smaller unit than the liter, so we use milliliters (mL)

1cc^3
cubic centimeter
=
milliliter

$$1 \text{ mL} = 10^{-3} \text{ L}$$

-or-

$$1000 \text{ mL} = 1 \text{ L}$$